Periodicity Chemistry Worksheet

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1. Which are metals? Circle your answers: C. Na. F. Cs. Ba, Ni.

Which metal in the list above has the most metallic character? Explain.

Cs - It is the metal with the † shielding & | ENC

Write the charge that each of the following atoms will have when it has a complete set of valence electrons forming an ion.

O 4 Na 1 F 1 N 3 Ca 2 Ar 6

What is the most common oxidation number for calcium? Explain.
 Ca¹² – Calcium has 2 valence electrons, it will lose 2 electrons.

Name two more elements with that oxidation number and explain your choice.

Any of the two: Be, Mg, Se, Ba, Ra – Same valence electrons / same family

- 4. What element in period 3 is a metalloid? Silicon
- When element with atomic number 118 is discovered, what family will it be in? Noble Gas.

Use the following word bank for questions 6 through 13

(Obviously, you will need to use several words more than once! ○)
Alkali metals Alkaline earth metals Halogens
Noble gases Transition metals Noble gases

- The __Alkali metals_____ have a single electron in the highest energy level.
- The __Alkaline earth metals__ achieve the electron configurations of noble gases by losing two electrons.
- The __transition metals__ are metals that can hold up to 10 electrons in their sublevel shape
- The _halogens ____ achieve the electron configuration of noble gases by gaining one electron.
- The <u>noble gases</u> have full s and p orbitals in the highest occupied energy levels.
- 11. The __noble gases_____ are stable and un-reactive.
- 12. The __halogens _____ are highly reactive and readily form salts with metals.
- 13. The __Alkaline earth metals are metals that are more reactive than the transition elements but less reactive than the alkali metals.
- 14. Predict the oxidation number based on the electron configuration shown.

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- 1. Understanding the eBook Periodic Worksheet Chemistry If8766 Answer
 - o The Rise of Digital Reading Periodic Worksheet Chemistry If8766 Answer
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Worksheet Chemistry If8766 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Worksheet Chemistry If8766 Answer
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Worksheet Chemistry If8766 Answer
 - Personalized Recommendations
 - Periodic Worksheet Chemistry If8766 Answer User Reviews and Ratings
 - Periodic Worksheet Chemistry If8766 Answer and Bestseller Lists
- 5. Accessing Periodic Worksheet Chemistry If8766 Answer Free and Paid eBooks
 - Periodic Worksheet Chemistry If8766 Answer Public Domain eBooks
 - Periodic Worksheet Chemistry If8766 Answer eBook Subscription Services
 - Periodic Worksheet Chemistry If8766 Answer Budget-Friendly Options

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 - o ePub, PDF, MOBI, and More
 - Periodic Worksheet Chemistry If8766 Answer Compatibility with Devices
 - Periodic Worksheet Chemistry If8766 Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Periodic Worksheet Chemistry If8766 Answer
 - Highlighting and Note-Taking Periodic Worksheet Chemistry If8766 Answer
 - Interactive Elements Periodic Worksheet Chemistry If8766 Answer
- 8. Staying Engaged with Periodic Worksheet Chemistry If8766 Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Worksheet Chemistry If8766 Answer
- 9. Balancing eBooks and Physical Books Periodic Worksheet Chemistry If8766 Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Worksheet Chemistry If8766 Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Worksheet Chemistry If8766 Answer
 - Setting Reading Goals Periodic Worksheet Chemistry If8766 Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Worksheet Chemistry If8766 Answer
 - Fact-Checking eBook Content of Periodic Worksheet Chemistry If8766 Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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