

Name _____ Date _____ Per _____

Periodic Trends Lab

Periodic law states that when the elements are arranged in order of increasing atomic number, there is a periodic repetition of their chemical and physical properties. Elements within the same group have the most similar chemical and physical properties, although not identical. Elements in neighboring groups may also have similar chemical and physical properties, but to a lesser degree. Trends in properties occur within groups and across periods.

PURPOSE: In this activity, you will

- Interpret group and periodic trends in atomic radius, electronegativity, and ionization energy
- Graph these trends

PROCEDURE:

1. Complete the pre-lab section
2. Using the data from **Table S** of your reference tables, prepare the graphs listed below. Staple your graphs to the back of this lab packet.

On each graph be sure to:

- **USE PENCIL!**
- Put an appropriate title and labels for each axis.
- Omit the transition elements.
- Connect the points you plot.
- Connect the points representing Group 1 elements with a green line.
- Connect the points representing Group 2 elements with a blue line.
- Connect the points representing Group 17 elements with a red line.
- Draw a triangle around the points representing the elements of Period 3.

GRAPH 1. Atomic radius vs. atomic number for elements 1-54.

GRAPH 2. Electronegativity vs. atomic number for elements 1-54.

GRAPH 3. Ionization energy vs. atomic number for elements 1-54.

3. Complete the conclusion section.

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Table of Contents Periodic Trends Pre Lab Answers

1. Understanding the eBook Periodic Trends Pre Lab Answers
 - The Rise of Digital Reading Periodic Trends Pre Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Pre Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Pre Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Pre Lab Answers
 - Personalized Recommendations
 - Periodic Trends Pre Lab Answers User Reviews and Ratings

- Periodic Trends Pre Lab Answers and Bestseller Lists
- 5. Accessing Periodic Trends Pre Lab Answers Free and Paid eBooks
 - Periodic Trends Pre Lab Answers Public Domain eBooks
 - Periodic Trends Pre Lab Answers eBook Subscription Services
 - Periodic Trends Pre Lab Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Pre Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Pre Lab Answers Compatibility with Devices
 - Periodic Trends Pre Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Pre Lab Answers
 - Highlighting and Note-Taking Periodic Trends Pre Lab Answers
 - Interactive Elements Periodic Trends Pre Lab Answers
- 8. Staying Engaged with Periodic Trends Pre Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Pre Lab Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Pre Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Pre Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Pre Lab Answers
 - Setting Reading Goals Periodic Trends Pre Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Pre Lab Answers
 - Fact-Checking eBook Content of Periodic Trends Pre Lab Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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