

## Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{1} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaCO}_3} \cdot \frac{56.08 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \Rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{1} \cdot \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 16.6 \text{ g Cu} \quad \text{TY}$$

$$\frac{9.65 \text{ g CuSO}_4}{1} \cdot \frac{1 \text{ mol CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.05\%}$$

3. If you start with 5.0 g of  $\text{C}_3\text{H}_8$  and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{1} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 8.2 \text{ g H}_2\text{O} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100$$

$$\frac{\text{PY} \cdot \text{TY}}{100} = \text{AY} = \frac{75 \cdot 8.2}{100} = \boxed{6.2 \text{ g H}_2\text{O}} \quad \text{AY}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{1} \cdot \frac{1 \text{ mol CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}} \quad \text{TY}$$

# Percent Yield Answer Key

**TD Snyder**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'TD Snyder' text.

## **Percent Yield Answer Key:**

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**OCR AS Chemistry Student Unit Guide** Mike Smith, 2009-06-26 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide Introduction includes advice on how to use the guide an explanation of the skills being tested by the assessment objectives an outline of the unit or module and depending on the unit

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**NTA NEET 40 Days Crash Course in Chemistry with 33 Online Test Series 3rd Edition** Disha Experts, 2018-12-17 This book contains an Access Code in the starting pages to access the 33 Online Tests NTA NEET 40 Days Crash Course in Chemistry is the thoroughly revised updated redesigned study material developed for quick revision

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