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# AP CHEMISTRY MULTIPLE CHOICE QUESTIONS Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed  
+ Doubled

(A) 95<sup>2</sup>

(B) 95

(C) (95)<sup>1/2</sup>

(D)  $\frac{1}{95}$

(E)  $\frac{1}{(95)^2}$



$C_2H_2(g)$  is reduced by  $H_2(g)$  in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of  $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

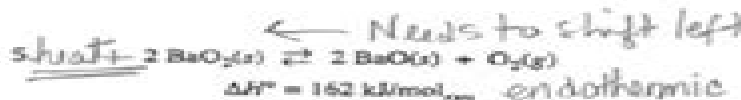
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains  $BaO_2(s)$  in equilibrium with  $BaO(s)$  and  $O_2(g)$  as represented by the equation above. Which of the following changes will increase the amount of  $BaO_2(s)$  in the vessel?

(A) Removing a small amount of  $O_2(g)$   $R \times$

(B) Removing a small amount of  $BaO(s)$   $X$

(C) Adding He gas to the vessel  $X$

(D) Lowering the temperature  $L$

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase  $P^\circ$



Shift towards fewer  
+ make gas



When 4.00 mol each of  $X(g)$  and  $Y(g)$  are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of  $Z(g)$  is produced. What is the value of the equilibrium constant,  $K_c$ ?

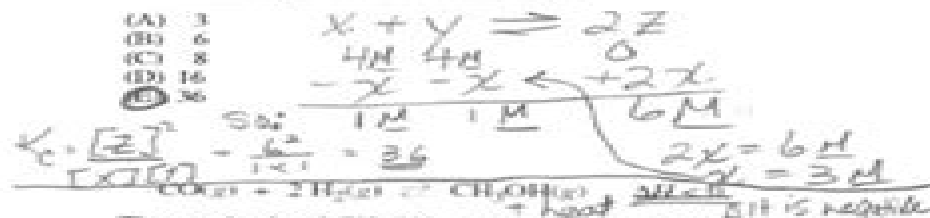
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of  $CH_3OH(g)$  from  $CO_2(g)$  and  $H_2(g)$  is represented by the equation above. The value of  $K_c$  for the reaction at 483 K is 14.5.

increase heat shifts left, decrease  $K_c$

Which of the following explains the effect on the equilibrium constant,  $K_c$ , when the temperature of the reaction system is increased to 650 K?

(A)  $K_c$  will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B)  $K_c$  will increase because there are more reactant molecules than product molecules.

(C)  $K_c$  will decrease because the reaction is exothermic.

(D)  $K_c$  is constant and will not change.

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**Clemens Wendtner**



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