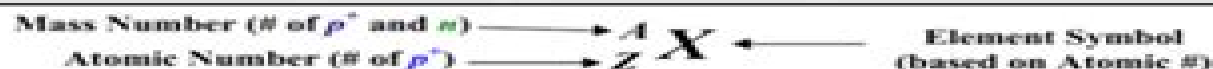


Chapter 21: Nuclear Chemistry21.1: The Nature of Nuclear Reactions

Nucleons: - the particles that make up a nucleus of an atom (protons, (${}^1_1p^+$ or 1_1H) and neutrons, (1_0n)).

Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Nuclide: a particular atom or isotope containing specific numbers of protons and neutrons



Radioactivity: - the particles and/or electromagnetic radiation that are emitted due to unstable nuclei.
 - all elements having atomic number 84 (Polonium) and greater are radioactive.

Nuclear Transmutation: - a process where radioactivity is resulted from the bombardment of nuclei by neutrons, protons or other nuclei.
 - in most cases, heavier elements are synthesized from lighter elements.

History of Radioactivity

- In 1896, **Wilhelm Roentgen** discovered X-ray by examining ray emitting from the outside of the cathode ray glass tube. It has the capability of passing through solid materials, but can be blocked by denser matters. It can also be exposed to photographic plate, resulting from an image of "seeing through" an container made of less dense material.
- In the same year, **Antoine Becquerel** discovered that uranium emits a ray onto a photographic plate in the absence of sunlight or other forms of energy.
- A few years later, **Marie and Pierre Curie** demonstrated that radiation can be emitted by other elements. They discovered two new elements (polonium and radium) based on their tendency to emit radiation (**radioactivity**).

Types of Radioactive Particle and Decay

- Alpha Particle (α particle):** - basically a helium nucleus (4_2He), commonly found during radioactive decay from heavier nuclide (the net result is to **increase the neutron to proton ratio** – more explanation in the next section).



- Beta Particle (β particle):** - basically an electron (${}^0_{-1}e$ or ${}^0_{-1}\beta$) that is emitted when the neutron to proton ratio is higher than the zone of stability (**a neutron is transformed to a proton and an electron as a result** – more explanation in the next section).
 - electrons have a mass number of 0 and an atomic number assignment of -1, due to its charge.



Review Nuclear Chemistry Modern Chemistry Answers

Wenbin Ji



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

Reviews of Modern Quantum Chemistry K. D. Sen, 2002 This important book collects together state of the art reviews of diverse topics covering almost all the major areas of modern quantum chemistry The current focus in the discipline of chemistry synthesis structure reactivity and dynamics is mainly on control A variety of essential computational tools at the disposal of chemists have emerged from recent studies in quantum chemistry The acceptance and application of these tools in the interfacial disciplines of the life and physical sciences continue to grow The new era of modern quantum chemistry throws up promising potentialities for further research Reviews of Modern Quantum Chemistry is a joint endeavor in which renowned scientists from leading universities and research laboratories spanning 22 countries present 59 in depth reviews Along with a personal introduction written by Professor Walter Kohn Nobel laureate Chemistry 1998 the articles celebrate the scientific contributions of Professor Robert G Parr on the occasion of his 80th birthday List of Contributors W Kohn M Levy R Pariser B R Judd E Lo B N Plakhutin A Savin P Politzer P Lane J S Murray A J Thakkar S R Gadre R F Nalewajski K Jug M Randic G Del Re U Kaldor E Eliav A Landau M Ehara M Ishida K Toyota H Nakatsuji G Maroulis A M Mebel S Mahapatra R Carb Dorca Nagy I A Howard N H March S B Liu R G Pearson N Watanabe S Ten no S Iwata Y Udagawa E Valderrama X Fradera I Silanes J M Ugalde R J Boyd E V Lude a V V Karasiev L Massa T Tsuneda K Hirao J M Tao J P Perdew O V Gritsenko M Gr ning E J Baerends F Aparicio J Garza A Cedillo M Galv n R Vargas E Engel A H ck R N Schmid R M Dreizler J Poater M Sol M Duran J Robles X Fradera P K Chattaraj A Poddar B Maiti A Cedillo S Guti rrez Oliva P Jaque A Toro Labb H Chermette P Boulet S Portmann P Fuentealba R Contreras P Geerlings F De Proft R Balawender D P Chong A Vela G Merino F Kootstra P L de Boeij R van Leeuwen J G Snijders N T Maitra K Burke H Appel E K U Gross M K Harbola H F Hameka C A Daul I Ciofini A Bencini S K Ghosh A Tachibana J M Cabrera Trujillo F Tenorio O Mayorga M Cases V Kumar Y Kawazoe A M K ster P Calaminici Z G mez U Reveles J A Alonso L M Molina M J L pez F Dugue A Ma anes C A Fahlstrom J A Nichols D A Dixon P A Derosa A G Zacarias J M Seminario D G Kanhere A Vichare S A Blundell Z Y Lu H Y Liu M Elstner W T Yang J Mu oz X Fradera M Orozco F J Luque P Tarakeshwar H M Lee K S Kim M Valiev E J Bylaska A Gramada J H Weare J Brickmann M Keil T E Exner M Hoffmann J Rychlewski

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U.S. Environmental Protection Agency Library System Book Catalog Holdings as of July 1973 United States. Environmental Protection Agency. Library Systems Branch,1974 **107-2 Hearing: The National Security Implications of The Strategic Offensive Reductions Treaty, S. Hrg. 107-806, July 25 and August 1, 2002, *,2003 **The National Security Implications of the Strategic Offensive Reductions Treaty** United States. Congress. Senate. Committee on Armed Services,2003**

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