

Calculate the partial pressure of nitrogen gas, the moles of dry gas present and the total pressure in the system. Show your work.

$$\begin{aligned}
 PV &= nRT \rightarrow P_{\text{H}_2} = nRT/V \rightarrow (0.0540 \text{ mol} \times 0.08206 \text{ L atm/K mol} \times 300 \text{ K})/2.0 \text{ L} = \\
 P_{\text{H}_2} &= 0.632 \text{ atm} \\
 &\rightarrow n_{\text{H}_2} = PV/RT \rightarrow (0.91 \text{ atm} \times 2.0 \text{ L})/(0.08206 \text{ L atm/K mol} \times 300 \text{ K}) = \\
 n_{\text{H}_2} &= 0.0720 \text{ mol} \\
 P_{\text{tot}} &= P_{\text{H}_2} + P_{\text{N}_2} \\
 &= 0.91 \text{ atm} + 0.632 \text{ atm} = P_{\text{tot}} = 1.54 \text{ atm}
 \end{aligned}$$

7. Calculate the volume of hydrogen gas produced at 4.50 atm and -30°C by the reaction of 26.5g of calcium with excess hydrobromic acid according to the unbalanced equation. Show your work.



& 244K

$$PV = nRT \rightarrow 26.5 \text{ g Ca} \times (1 \text{ mol Ca}/40.08 \text{ g Ca}) \times (1 \text{ mol H}_2/1 \text{ mol Ca}) = 0.661 \text{ mol H}_2$$

$$V = nRT/P = (0.661 \text{ mol} \times 0.08206 \text{ L atm/K mol} \times 244 \text{ K})/4.50 \text{ atm} = V = 2.94 \text{ L}$$

8. Hydrogen Sulfide reacts with sulfur dioxide to produce water and sulfur. If 3.2g of sulfur was produced under the following conditions: 6.0 L @ 750 torr, what is the temperature of the hydrogen sulfide gas, in °C?



$$3.2 \text{ g S} \times (1 \text{ mol S}/32.07 \text{ g S}) \times (2 \text{ mol H}_2\text{S}/3 \text{ mol S}) = 0.067 \text{ mol H}_2\text{S}$$

$$PV = nRT \rightarrow T = PV/nR \rightarrow (0.987 \text{ atm} \times 6.0 \text{ L})/(0.067 \text{ mol} \times 0.08206 \text{ L atm/K mol}) = 1077.12 \text{ K}$$

$$= 804^\circ\text{C}$$

SL&B - Student will apply Kinetic Molecular Theory when describing a novel, real-world situation in terms of collisions and motions of particles.

9. In the summer, the tire pressure sensor in your car alerts you that your tires need to be checked. At the gas station, you find that the tire pressure is too high inside the tire. However, you have not added additional gas. Using the Kinetic Molecular Theory, explain why this occurs, making sure to refer to the motion of the gas molecules.

Since the summer is known for high temps, a high T would cause for the motion of gas molecules inside the tire to increase. With the increase in motion comes an increase in the number of collisions. With more collisions occurring, this results in an increase in pressure inside the tire.

Review Gases Key

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Agricultural News -- A Fortnightly Review of the Imperial Department of Agriculture for the West Indies -- Volume 1., 1902

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Digest; Review of Reviews Incorporating Literary Digest, 1928

Soft Computing in Condition Monitoring and Diagnostics of Electrical and Mechanical Systems Hasmat Malik, Atif Iqbal, Amit Kumar Yadav, 2020-01-17 This book addresses a range of complex issues associated with condition monitoring CM fault diagnosis and detection FDD in smart buildings wide area monitoring WAM wind energy conversion systems WECSs photovoltaic PV systems structures electrical systems mechanical systems smart grids etc The book's goal is to develop and combine all advanced nonintrusive CMFD approaches on a common platform To do so it explores the main components of various systems used for CMFD purposes The content is divided into three main parts the first of which provides a brief introduction before focusing on the state of the art and major research gaps in the area of CMFD The second part covers the step by step implementation of novel soft computing applications in CMFD for electrical and mechanical systems In the third and final part the simulation codes for each chapter are included in an extensive appendix to support newcomers to the field

Journal of Research of the National Bureau of Standards United States. National Bureau of Standards, 1976

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Table of Contents Review Gases Key

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 - The Rise of Digital Reading Review Gases Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Gases Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Gases Key
 - User-Friendly Interface
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 - Review Gases Key Enhanced eBook Features
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 - Distinguishing Credible Sources

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