

How to Identify the Limiting & Excess Reactants/Reagents

Molar Mass:

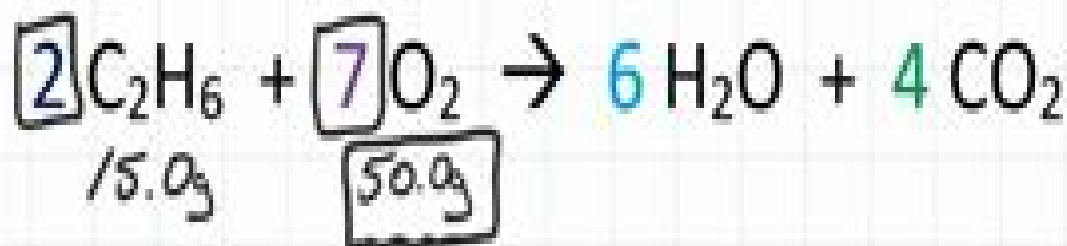
✖ $C_2H_6 = 30.08 \text{ g/mol}$

✖ $O_2 = 32.00 \text{ g/mol}$

$H_2O = 18.02 \text{ g/mol}$

1.) Calculate how much of one reactant will react with the other reactant

Ex. What is the limiting reactant when 15.0 g of C_2H_6 react with 50.0 g of O_2 ?



$$\overset{A}{\underset{15.0g}{\cancel{C_2H_6}}} \left(\frac{1 \cancel{\text{mol } C_2H_6}}{30.08g \cancel{C_2H_6}} \right) \left(\frac{7 \cancel{\text{mol } O_2}}{2 \cancel{\text{mol } C_2H_6}} \right) \left(\frac{32.00g \cancel{O_2}}{1 \cancel{\text{mol } O_2}} \right) = \underline{\underline{55.9g O_2}} \quad \overset{B}{}$$

$O_2 = \text{L.R.}$

$$\overset{A}{\underset{50.0g}{\cancel{O_2}}} \left(\frac{1 \cancel{\text{mol } O_2}}{32.00g \cancel{O_2}} \right) \left(\frac{2 \cancel{\text{mol } C_2H_6}}{7 \cancel{\text{mol } O_2}} \right) \left(\frac{30.08g \cancel{C_2H_6}}{1 \cancel{\text{mol } C_2H_6}} \right) = 13.4g C_2H_6 \quad \overset{B}{}$$

$C_2H_6 = \text{E.R.}$

If reacting all of Reactant A requires **more** of Reactant B than is available, then **B** is the Limiting Reactant.

If reacting all of Reactant A requires **less** of Reactant B than is available, then **A** is the Limiting Reactant.

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