

Gas Behavior in the Real World

Name _____

Skill Review Skill – Student will accurately describe the relationship between P, V, T and n of gases.)

In your everyday life, you often experience the different gas laws discussed in class, in action. In the table below are some situations that may be familiar to you. Think about each situation and in column 2, describe in your own words what is happening on the microscopic (molecular) level. In column 3, write the name of the simple law (Avogadro's, Boyle's, Charles' or Dalton's) that best describes the situation. In the two extra rows at the bottom, add some additional situations, not already discussed, that you have experienced.

Everyday Situation	Microscopic Description	Law
1. The pressure shown on a tire gauge doubles as the mass of air is doubled (at a constant temperature).		Avogadro's
2. A balloon over the mouth of a bottle containing air begins to expand when it is moved in to sunlight.		Charles'
3. An car piston compresses gases while the engine is working.		Boyle's
4. An inflated rubber raft gets softer when some of the gas is allowed to escape.		Avogadro's
5. A balloon placed in a sauna increases in size.		Charles'
6. A hot air balloon lands when burners under the balloon's open end turn off.		Charles'
7. When you squeeze an inflated balloon, it seems to push back harder in to your hand.		Boyle's
8. A highly pressurized tank of helium gas will fill hundreds of balloons.		Boyle's
9. Red, yellow and blue bingo balls are flying around inside a bingo caller's dispenser. The effect on the sides of the container is the same as if an equal number of all red bingo balls are in the container.		Dalton's
10. When the number of bingo balls is doubled, the effect on the sides of the container is doubled.		Avogadro's
11.		

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