

## Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be  $\text{CO}_2$  and  $\text{H}_2\text{O}$  (assuming a complete reaction).



Note that Mg can only form  $\text{Mg}^{2+}$  and I can only form  $\text{I}^-$ , so the product will must have a 1:2 cation:anion ratio.



Note that the product is not  $\text{H}_2\text{Cl}_2$ . It is important to recognize that  $\text{CuCl}_2$  is made of three ions,  $\text{Cu}^{2+}$  and two  $\text{Cl}^-$ .



In this question, you must recognize that perchlorate,  $\text{ClO}_4^-$ , and hydroxide,  $\text{OH}^-$ , are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.



When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to  $\text{CO}_2$  and a metal oxide.



Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants ( $\text{H}^+$  in this case), will end up in its elemental form in the products ( $\text{H}_2$ ) and the species that IS in its elemental form (Zn) will end up ionized ( $\text{Zn}^{2+}$ ). Note that zinc can *only* form a  $\text{Zn}^{2+}$  ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is  $\text{H}_2$ , not H.



Again notice that one species is in its elemental form (Na). The magnesium in  $\text{MgCl}_2$  is an ion ( $\text{Mg}^{2+}$ ), but is transformed into its elemental state (Mg), while the Na is converted into an ion ( $\text{Na}^+$ , sodium *only* forms a +1 ion).



Recognize that carbonate is a polyatomic ion ( $\text{CO}_3^{2-}$ ) and that the cations are already stable ions ( $\text{Ca}^{2+}$  and  $\text{K}^+$ ).



Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product,  $\text{K}^+$  and  $\text{Cl}^-$  are formed. Note that we would NOT write the product as  $\text{K}_2\text{Cl}$ . Ionic compounds are written in their simplest forms.



Note that phosphate ( $\text{PO}_4^{3-}$ ) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of  $\text{Ba}:\text{PO}_4$  in order to balance the charge.



Notice that sulfate ( $\text{SO}_4^{2-}$ ) is a polyatomic ion and that potassium exists as a +1 ion ( $\text{K}^+$ ). Also, this is an acid-base reaction, so the products should be salt and water.



Another decomposition that will generate  $\text{CO}_2$  and a metal oxide (note that Al forms a +3 ion,  $\text{Al}^{3+}$ , and monatomic oxygen will have a -2 charge,  $\text{O}^{2-}$ ).

# Predicting Products Answer Key

**AW Chickering**



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