

SPECTRUM TWO



User's Guide



Perkin Elmer Ftir Manual

Daniela Niemeyer



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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1220 \text{ J}$
 $550 \text{ J} = \Delta h \cdot q_{\text{water}} = 550 \text{ J}$

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 for water $4.184 \text{ J/g} \cdot \text{C}$ 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by 8.1 C when 207 J of heat is added 2

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 $14.7 \text{ J} = 3.14 \Delta H_{\text{rxn}} = n \cdot q_{\text{rxn}} = n \cdot q_{\text{calorimeter}} = m \cdot c \cdot \Delta T$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of
 25.0 C and heat it to 145.0 C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the
 process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 C

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 of 1.00 M NaOH at 25.0 C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 C calculate the
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web 6 a 218 g sample of steam at 121 $^{\circ}\text{C}$ is cooled to ice at 14 $^{\circ}\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1 41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 $\times 10^4 \text{ J}$ 2 6 23 $\times 10^4 \text{ J}$ 3 1 71 $\times 10^4 \text{ J}$ 4 1 11 $\times 10^5 \text{ J}$ 5 2 11 $\times 10^5 \text{ J}$ 6

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