Period Date Tuesday, November 1 2016 Name

Periodic Trends

ATOMIC RADIUS

What trend in atomic radius do you see as you go down a group/family on the periodic table?

2. What causes this trend?

3. What trend in atomic radius do you see as you go across a period/row on the periodic table?

4. What causes this trend?

Circle the atom in each pair that has the largest atomic radius.

a) Al B

b) S O c) Br Cl

d) Na Al e) O F f) Mg Ca

Put the following elements in order from smallest to largest atomic radius and explain why: C. O. Sn. Sr.

ELECTRONEGATIVITY

Define electronegativity

8. How does the ionic radius of a nonmetal compare with its atomic radius?

9. What trend in electronegativity do you see as you go down a group/family on the periodic table?

10. What causes this trend?

11. What trend in electronegativity do you see as you go across a period/row on the periodic table?

12. What causes this trend?

Circle the atom in each pair that has the greater electronegativity.

a) Ca Ga b) Li O c) Cl S d) Br As

e) Ba Sr

f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions?

15. Which group tends to form +2 ions?

16. Which group tends to form -1 ions?

17. Which group tends not to form ions or react? _____

Periodic Trends Quiz Answer Key

M Tight

Periodic Trends Quiz Answer Key:

Periodic Trends Quiz Answer Key Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Periodic Trends Quiz Answer Key**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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