

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O C **Be** **Ne** Same Period
 - Na Rb **Fr** **H** Same Group
 - Pb** **C** Sn Si Same Group
 - Au W S **Fr** **Ne** Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O C **Be** **Ne** Same Period
- Na Rb **Fr** **H** Same Group
- Pb** **C** Sn Si Same Group
- Au W S **Fr** **Ne** Zn Challenge

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Jianjun Gao



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Scientific and Technical Information Database **Homenatge professor Josep M. Costa. Trends in electrochemistry and corrosion at the beginning of the 21st century** Pere-Lluís Cabot Julià, 2004-04-13 Este libro est dedicado al Profesor Josep M Costa en ocasi n de su 70 aniversario Re ne un total de 73 art culos y revisiones originales tanto cient ficas como tecnol gicas escritas en espa ol e ingl s por unos 250 investigadores de todo el mundo y que son exponentes representativos de la investigaci n internacional en materias de gran inter s en la Electroqu mica y la Corrosi n de principios de este siglo XXI El libro se ha estructurado en dos grandes secciones La primera secci n correspondiente a la Electroqu mica consta de 33 trabajos distribuidos en 5 cap tulos dedicados a los campos de Electroqu mica Molecular Electrodeposici n Electroodos Modificados Descontaminaci n Electroqu mica y Sensores y Electroan lisis La segunda secci n relativa a la Corrosi n comprende 40 trabajos que se agrupan en otros 5 cap tulos que versan sobre Corrosi n en Ambientes Corrosivos Seleccionados Protecci n contra la Corrosi n y Monitorizaci n Recubrimientos Nuevos Materiales y Tratamientos y Educaci n en la Corrosi n This book is dedicated to Professor Josep M Costa in occasion of his 70th birthday It collects a total number of 73 original articles and reviews both scientific and technologic written in English and Spanish by about 250 researchers of all around the world who are representative exponents of the international research in topics of great interest in Electrochemistry and Corrosion at the beginning of the 21st Century The book has been structured in two large sections The first section corresponds to Electrochemistry and includes 33 articles distributed into five chapters related to the fields of Molecular Electrochemistry Electrodeposition Modified Electrodes Electrochemical Depollution and Sensors and Electroanalysis The second section is related to Corrosion and contains 40 articles gathered into other five chapters devoted to Corrosion in Selected Environments Corrosion Protection and Monitoring Coatings New Materials and Treatments and Corrosion Education **Shriver and Atkins' Inorganic Chemistry** Peter Atkins, 2010 Inorganic Chemistry fifth edition represents an integral part of a student s chemistry education Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text culminating in the cutting edge research topics of the Frontiers which illustrate the dynamic nature of inorganic chemistry **Surveyor and Municipal and County Engineer** ,1930 **Cumulated Index Medicus** ,1985 **Physics Briefs** ,1990 **The Surveyor & Municipal & County Engineer** ,1930 **Congressional Record** United States. Congress, 1973 Franciscan Almanac ,1968 **National Catholic Almanac** ,1968 *Science* John Michels (Journalist), 2007 A weekly record of scientific progress **Mathematical Reviews** ,2005 **St. Anthony's Almanac** ,1968 **New Scientist** ,1989 **Metals Abstracts** ,1999 The Geographical Magazine Michael Huxley, 1973 Vols for 19 include a separate section called GM news and reviews **Congressional Record** United States. Congress, 1972 **District Fifty News** United Mine Workers of America. District 50, 1956

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