

Jayden Zayas

Class of 2023

"You have brains in your head. You have feet in your shoes. You can steer yourself in any direction you choose." – Dr. Seuss

We are so proud of you and cannot wait to see what your future holds!
Love you,
Mom, Dad
& Jeremy



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Half
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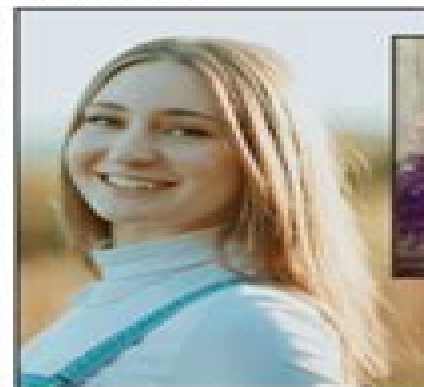
Harper Sorrells

Harper,
May you always
dance like no one
is watching!

We love you!
Mom, Dad, and
Hazey

Isabella Boston

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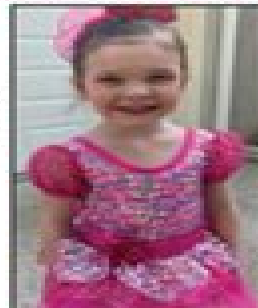
It has been an absolute joy watching you dance these past 12 years!
Love always,
Mom, Dad, Ethan, Casey, the boys, Alex, Mackenzie, Jake

Quarter
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Norah Reagan Turner

We are so proud of
you, beautiful girl!

Love,
Mommy, Daddy
& Liam



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web google classroom for the following redox equation $\text{MnO}_4^{-} + \text{aq} + \text{SO}_2 \rightarrow \text{Mn}^{2+} + \text{aq} + \text{HSO}_4^{-}$ what will be the stoichiometric coefficient of HSO_4^{-} if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} \rightarrow 2\text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h ions or the h₂o pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate $\text{Cr}_2\text{O}_7^{2-}$ and chloride cl ions in acidic solution created by jay

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reaction in acid solution using the half reaction method 8 h₂o₂ cr₂o₇²⁻ 2 9 teo₃ 2 n₂o₄ o₂ cr₃ te 10 reo₄ 11 pbo₂ io i₂ pb₂ 12 as no₃ io₃ re io₃ h₂aso₄

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h⁺ 2no²⁻ 2 c 6 h 6 o 6 2h⁺ 2 o²⁻ 2no choose 1 answer c 6 h 8 o 6 and no²⁻ a c 6 h 8 o 6 and no²⁻ no 2 and c 6 h 8 o 6 b no²⁻ and c 6 h 8 o 6 h⁺ and no²⁻ c h⁺ and

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