



Republic of the Philippines  
**PANGASINAN STATE UNIVERSITY**  
URDANETA CAMPUS  
Urdaneta City, Pangasinan

# **On the Job Training NARRATIVE REPORT**

Submitted to:

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Student Trainee

May 2013

# Sample Ojt Narrative Report For Students

**United States. Office of Management  
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web google classroom for the following redox equation  $\text{MnO}_4^{-}$  aq  $\text{SO}_2$  g  $\text{Mn}^{2+}$  aq  $\text{HSO}_4^{-}$  aq what will be the stoichiometric coefficient of  $\text{HSO}_4^{-}$  a if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h2o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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