Chapter 10 Nuclear Chemistry

- How is the product isotope different from the reactant isotope in alpha decay? The product isotope has two fewer protons and two fewer neutrons than the reactant isotope.
- Circle the letters that identify each event that takes place during beta decay.
 - a. A proton decomposes into a neutron and an electron.
 - (b) A neutron decomposes into a proton and an electron.
 - c.) An electron is emitted from the nucleus.
 - A neutron is emitted from the nucleus.
- Why are beta particles more penetrating than alpha particles?Beta particles have a smaller mass and a faster speed.
- 10. Is the following sentence true or false? All nuclear radiation consists of charged particles.
- What is a gamma ray? A gamma ray is a penetrating ray of energy emitted by an unstable nucleus.
- 12. How fast do gamma rays travel through space? Gamma rays travel through space at the speed of light.
- 13. Complete the following table about nuclear radiation.

Characteristics of Nuclear Radiation			
Radiation Type	Charge	Mass (amu)	Usually Stopped By
Alpha particle	-1	4	Paper or clothing
Beta particle	1-	1 1836	Aluminum sheet
Garmena raiy	0	0	Several meters of concrete

Effects of Nuclear Radiation (pages 296-297)

- 14. How does nuclear radiation affect atoms? Nuclear radiation can ionize atoms.
- Is the following sentence true or false? One potential danger of radon gas is that prolonged exposure to it can lead to lung cancer.

Detecting Nuclear Radiation (page 297)

- Name two devices that are used to detect nuclear radiation.
 - а, Geiger counters b, Film badges

Review Nuclear Chemistry Section 2 Answer Key

LL Leslie

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