

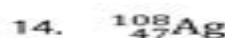
Unit 3 Study Guide: Atomic Structure and Nuclear Chemistry

Section 1: Atomic Structure

Define the following terms.

1. atom
2. Why is an atom neutral?
3. electron cloud
4. electrons
5. What determines which element an atom is?
6. neutrons
7. Which equals one amu—the mass of one atom of carbon-12, or 1/12 the mass of one atom of carbon-12?
8. atomic number
9. mass number
10. hyphen notation

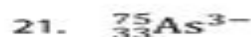
Give the number of protons, neutrons, and electrons for the following:



Section 2: Ions and Isotopes

15. Describe how cations are formed.
16. Describe how anions are formed.
17. What type of element generally forms cations?
18. How are isotopes of the same element alike and how are they different?

Give the number of protons, neutrons, and electrons for the following:



Section 3: Average Atomic Mass

23. Chlorine has two naturally occurring isotopes, Cl-35 and Cl-37. Using the periodic table, determine which of these two isotopes of chlorine is more abundant.

Review Nuclear Chemistry Section 1 Answers

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