

CHAPTER 17

RADIOACTIVITY AND NUCLEAR CHEMISTRY

Practice Problems

1. (a) Write a nuclear equation for alpha decay of Po-210 .



- (b) Write a nuclear equation for the beta decay of Th-231 .



- (c) Write a nuclear equation for positron emission by Co-57 .



2. A 10.0-g sample of a radioactive nuclide is administered to a patient for cancer diagnosis. If the nuclide has a half-life of 1.5 hours, how much of the nuclide remains in the patient after 4.5 days?

$$\text{4.5 days} \times \frac{24 \text{ hours}}{1 \text{ day}} \times \frac{60 \text{ min}}{1 \text{ hour}} = 6480 \text{ min} \quad \frac{6480 \text{ min}}{90 \text{ min}} = 72 \text{ half-lives}$$

- (a) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (b) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (c) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (d) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (e) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (f) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (g) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (h) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (i) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (j) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$

3. A mammoth skeleton has a carbon-14 content of 1.56% of that found in living organisms. When did the mammoth live?

The age of the mammoth is 14,400 years. The half-life of C-14 is 5,730 years.

$$\frac{1.56\%}{100\%} = \left(\frac{1}{2}\right)^n \quad n = 4 \quad 4 \times 5,730 \text{ years} = 22,920 \text{ years}$$

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