

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

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JS Bruner



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Table of Contents Review Gases Section 3 Modern Chemistry Answers

1. Understanding the eBook Review Gases Section 3 Modern Chemistry Answers
 - The Rise of Digital Reading Review Gases Section 3 Modern Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Gases Section 3 Modern Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Gases Section 3 Modern Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Gases Section 3 Modern Chemistry Answers
 - Personalized Recommendations
 - Review Gases Section 3 Modern Chemistry Answers User Reviews and Ratings

- Review Gases Section 3 Modern Chemistry Answers and Bestseller Lists
- 5. Accessing Review Gases Section 3 Modern Chemistry Answers Free and Paid eBooks
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 - ePub, PDF, MOBI, and More
 - Review Gases Section 3 Modern Chemistry Answers Compatibility with Devices
 - Review Gases Section 3 Modern Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Gases Section 3 Modern Chemistry Answers
 - Highlighting and Note-Taking Review Gases Section 3 Modern Chemistry Answers
 - Interactive Elements Review Gases Section 3 Modern Chemistry Answers
- 8. Staying Engaged with Review Gases Section 3 Modern Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Gases Section 3 Modern Chemistry Answers
- 9. Balancing eBooks and Physical Books Review Gases Section 3 Modern Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Gases Section 3 Modern Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Gases Section 3 Modern Chemistry Answers
 - Setting Reading Goals Review Gases Section 3 Modern Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Gases Section 3 Modern Chemistry Answers
 - Fact-Checking eBook Content of Review Gases Section 3 Modern Chemistry Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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