

# Acids & Bases: Assessment Review

Name \_\_\_\_\_

*Skill – The student will be able to identify the acid, base, conjugate acid, and conjugate base in an acid-base reaction.*

1. Write a chemical equation showing how  $\text{HCO}_3^-$  behaves as an acid when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

2. Write a chemical equation showing how ammonia ( $\text{NH}_3$ ) behaves as a base when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

*Skill – The student will be able to calculate with the  $[\text{H}^+]$  or  $[\text{OH}^-]$  from the ion product ( $K_w$ ) or pH from pOH (a one-step pH block calculation).*

3. Recreate the calculation flowchart on the diagram provided. (Be careful...it's been flipped!)



4. What is  $[\text{OH}^-]$  if a solution has  $[\text{H}^+] = 7.1 \times 10^{-4} \text{ M}$ ? Is the solution acidic, basic, or neutral?

5. Calculate  $[\text{H}^+]$  in a solution if  $[\text{OH}^-] = 5.1 \times 10^{-11} \text{ M}$ . Is the solution acidic, basic, or neutral?

6. Calculate the  $[\text{H}^+]$  in a solution that shows a pOH of 11.75. Is the solution acidic, basic, or neutral?

7. What is the pOH of a solution if  $[\text{OH}^-] = 5.0 \times 10^{-4} \text{ M}$ ? Is the solution acidic, basic, or neutral?

# Review Acids And Bases Answer Key Holt

**R Bogdan**



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