

Lesson 2. Nuclear Chemistry & Energy

Outline

- 2.1 Nuclear Reactions and Radioactivity
- 2.2 Nuclear Stability
- 2.3 Kinetics of Radioactive Decay
- 2.4 Energetics of Nuclear Reactions
- 2.5 Fusion, Transmutation, Fission
- 2.6 Interaction of Radiation and Matter

Nuclear Chemistry

Nuclear chemistry is a study of nuclear reactions. Nuclear reactions are changes in matter originating in the nucleus of an atom. When a nuclei change spontaneously, it emits radiation and that said is for radioactivity. Radioactive elements are widely used in medicine: diagnostic tools and as a means of treatment especially for cancer. They are also used as test substances to determine the mechanisms of chemical reactions, to trace the movement of atoms in biological systems, and as date important historical artifacts. Nuclear reactions are also used both to generate electricity and to create weapons of massive destruction.

This lesson discussed nuclear reactions and its applications.

2.1 Nuclear Reactions and Radioactivity

Nuclear Reactions

One of the nuclear reactions in the universe is the formation of carbon-14, which is a result of collision of gas molecules in the atmosphere with cosmic rays. Cosmic rays are particles travelling at high speeds. Majority of cosmic rays are atomic nuclei. Roughly, 80% of cosmic rays are hydrogen nuclei and about 10% are helium nuclei. Some cosmic rays emanate from the sun and others originate outside the solar system. The energies of cosmic rays are far higher than chemical energies. Chemical energies are usually expressed in kilojoules per mole (kJ/mol) but all cosmic rays are expressed in electron-volts (eV). One eV is equal to (1.6×10^{-19}) kJ/mol.

As cosmic rays enter the atmosphere, it collided with gas molecules. These collisions initiate nuclear reactions in the atmosphere. One of this nuclear reactions leads to the formation of the radioactive carbon-14. Carbon-14 is unstable isotope of carbon. It undergoes spontaneous radioactive decay or disintegration, emitting particles from the nucleus. Carbon-14 is formed when a free neutron is absorbed by nitrogen nucleus resulting to carbon nucleus and proton. Nitrogen nucleus and neutron are the reactants and carbon nucleus and proton are the products. The nuclear reaction is summarized



Note that the reactants and products of a nuclear reaction are summarized above as atomic notation, ${}^A_Z\text{X}$, where **X** is the symbol of the atom, **A** is the mass number and **Z** is the atomic number or nuclear charge. Similar notation is used for subatomic particles, neutrons (${}^1_0\text{n}$), protons (${}^1_1\text{p}$), and electrons (${}^0_{-1}\text{e}$).

Nuclear reactions also obey conservation rules, hence, nuclear reactions are balanced. Balancing nuclear equation is with respect to both mass number and nuclear charge. For the nuclear equation to form carbon-14, the sum of the mass numbers for each side of the equation is 14, and the sum of the nuclear charges on each side of the equation is 7, therefore the nuclear equation is balanced.

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Arturo Cuomo



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International Review of Cytology, 1984-01-04

Energy Research Abstracts, 1993

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Radiopharmaceuticals and Other Compounds Labelled with Short-Lived Radionuclides Michael J. Welch, 2013-09-24. Radiopharmaceuticals and Other Compounds Labelled with Short-Lived Radionuclides covers through both review and contributed articles the potential applications and developments in labeling with short-lived radionuclides whose use is restricted to institutions with accelerators. The book discusses the current and potential use of generator-produced radionuclides as well as other short-lived radionuclides and the problems of quality control of such labeled compounds. The book is useful to nuclear medicine physicians.

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Table of Contents Review Section 2 Nuclear Chemistry

1. Understanding the eBook Review Section 2 Nuclear Chemistry
 - The Rise of Digital Reading Review Section 2 Nuclear Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Section 2 Nuclear Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Section 2 Nuclear Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Section 2 Nuclear Chemistry
 - Personalized Recommendations
 - Review Section 2 Nuclear Chemistry User Reviews and Ratings
 - Review Section 2 Nuclear Chemistry and Bestseller Lists
5. Accessing Review Section 2 Nuclear Chemistry Free and Paid eBooks
 - Review Section 2 Nuclear Chemistry Public Domain eBooks
 - Review Section 2 Nuclear Chemistry eBook Subscription Services
 - Review Section 2 Nuclear Chemistry Budget-Friendly Options
6. Navigating Review Section 2 Nuclear Chemistry eBook Formats

- ePub, PDF, MOBI, and More
 - Review Section 2 Nuclear Chemistry Compatibility with Devices
 - Review Section 2 Nuclear Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Section 2 Nuclear Chemistry
 - Highlighting and Note-Taking Review Section 2 Nuclear Chemistry
 - Interactive Elements Review Section 2 Nuclear Chemistry
 8. Staying Engaged with Review Section 2 Nuclear Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Section 2 Nuclear Chemistry
 9. Balancing eBooks and Physical Books Review Section 2 Nuclear Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Section 2 Nuclear Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Review Section 2 Nuclear Chemistry
 - Setting Reading Goals Review Section 2 Nuclear Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Review Section 2 Nuclear Chemistry
 - Fact-Checking eBook Content of Review Section 2 Nuclear Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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