

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. a A chemical bond between atoms results from the attraction between the valence electrons and _____ of different atoms.
(a) nuclei (c) isotopes
(b) inner electrons (d) Lewis structures
2. b A covalent bond consists of _____
(a) a shared electron. (c) two different ions.
(b) a shared electron pair. (d) an octet of electrons.
3. a If two covalently bonded atoms are identical, the bond is identified as _____
(a) nonpolar covalent. (c) ionic.
(b) polar covalent. (d) dipolar.
4. b A covalent bond in which there is an unequal attraction for the shared electrons is _____
(a) nonpolar. (c) ionic.
(b) polar. (d) dipolar.
5. c Atoms with a strong attraction for electrons they share with another atom exhibit _____
(a) zero electronegativity. (c) high electronegativity.
(b) low electronegativity. (d) Lewis electronegativity.
6. c Bonds that possess between 5% and 50% ionic character are considered to be _____
(a) ionic. (c) polar covalent.
(b) pure covalent. (d) nonpolar covalent.
7. a The greater the electronegativity difference between two atoms bonded together, the greater the bond's percentage of _____
(a) ionic character. (c) metallic character.
(b) nonpolar character. (d) electron sharing.
8. The electrons involved in the formation of a chemical bond are called _____
valence electrons
9. A chemical bond that results from the electrostatic attraction between positive and negative ions is called a(n) _____
ionic bond

Review Chemical Bonding Section 1 Answers

S Marginson



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