



## Unit: Chemical Bonding Answer Key "Covalent Bonding" - WS #3

**Directions:** Please answer each question using your notes from today's class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

*A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example:  $C_3H_8$*

*A structural formula uses lines instead of dots to show shared pairs of electrons.*

*Example:  $H-H$*

*An electron dot formula uses dots instead of lines to show shared pairs of electrons.*

*Example:  $H:H$*

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

*The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.*

3. Draw the electron dot and structural formulas for the following molecules.

| Substance | Rough copy (work!) | Electron dot                                                                              | Structural                                              |
|-----------|--------------------|-------------------------------------------------------------------------------------------|---------------------------------------------------------|
| a) $Cl_2$ |                    | <i>Note: The electron dot structures have dots in between each atom instead of lines.</i> | $:\ddot{Cl}-\ddot{Cl}:$                                 |
| b) $H_2S$ |                    |                                                                                           | $\begin{array}{c} :\ddot{S}: \\   \\ H \end{array} - H$ |
| c) $BrI$  |                    |                                                                                           | $:\ddot{I}-\ddot{Br}:$                                  |

# Review Answers For Covalent Bonding

**Joacim Rocklöv**



## **Review Answers For Covalent Bonding:**

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