

Chapter 7 Chemical Reactions

Section 7.2 Types of Reactions

(pages 199–205)

This section discusses how chemical reactions are classified into different types.

Reading Strategy (page 199)

Previewing. Skim the section and begin a concept map like the one below that identifies types of reactions with a general form.

As you read, add the general forms of each type of reaction. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.



Classifying Reactions (pages 199–204)

1. Name five general types of chemical reactions. Synthesis, decomposition, single replacement, double replacement, and combustion.
2. Circle the letter of each equation that represents a synthesis reaction.
 - a. $2\text{Na} + \text{Cl}_2 \longrightarrow 2\text{NaCl}$
 - b. $2\text{NaCl} \longrightarrow 2\text{Na} + \text{Cl}_2$
 - c. $2\text{H}_2\text{O} \longrightarrow 2\text{H}_2 + \text{O}_2$
 - d. $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$
3. Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction. true
4. Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide. $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
5. Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
 - a. copper(II) nitride and silver oxide
 - b. copper(II) nitrate and silver
 - c. copper(II) oxide and silver nitrate
 - d. copper, nitrogen, and silver oxide

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Table of Contents Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets

1. Understanding the eBook Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets and Bestseller Lists
- 5. Accessing Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Public Domain eBooks
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets eBook Subscription Services
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Budget-Friendly Options
- 6. Navigating Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Compatibility with Devices
 - Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Interactive Elements Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- 8. Staying Engaged with Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Setting Reading Goals Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets

- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 1chemical Reactions Worksheets
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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