

Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: $[\text{Ne}] 3s^2 3p^2$ *may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an $s^2 p^3$ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a $4s^2 4p^5$ electron configuration **Br, bromine**
- the 4th period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$ **O, oxygen**
- $4s^2$ **Ca, Calcium**
- $4s^2 3d^{10}$ **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

Prentice Hall Chemistry 3 Practice Problems Answers

Arturo Cuomo



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Introduction to Chemical Principles Howard Stephen Stoker, 1999 Chemistry for students who need full exposure to general chemistry but in compact one semester 17 chapter paperback format Strong emphasis on problem solving with over 5000 problems in end of chapter material arranged in matched pairs More real life applications added to this edition plus faces of chemistry **High School Chemistry Teachers Magazine** ,1967

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