

Name _____ Date _____ Class _____

17

THERMOCHEMISTRY

SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK (pages 505–510)

This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat.

► Energy Transformations (page 505)

- What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions? thermochemistry
- Where the use of energy is concerned (in a scientific sense), when is work done?
Work is done when a force is used to move an object.
- Circle the letter next to each sentence that is true about energy:
 - ☒ Energy is the capacity for doing work or supplying heat.
 - ☒ Energy is detected only because of its effects.
 - ☐ Heat is energy that transfers from one object to another because they are at the same temperature.
 - ☒ Gasoline contains a significant amount of chemical potential energy.
- Circle the letter next to each sentence that is true about heat.
 - ☒ One effect of adding heat to a substance is an increase in the temperature of that substance.
 - ☐ Heat always flows from a cooler object to a warmer object.
 - ☒ If two objects remain in contact, heat will flow from the warmer object to the cooler object until the temperature of both objects is the same.

► Exothermic and Endothermic Processes (pages 506–507)

- What can be considered the “system” and what are the “surroundings” when studying a mixture of chemicals undergoing a reaction? Write your answers where indicated below.

System: The mixture of chemicals itself is considered the system.

Surroundings: Everything but the mixture of chemicals is the surroundings, but practically speaking, the immediate vicinity of the system.

Prentice Hall Chemistry Chapter 1 assessment Answers

Charles R. Cantor, Paul R. Schimmel



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Table of Contents Prentice Hall Chemistry Chapter 1assessment Answers

1. Understanding the eBook Prentice Hall Chemistry Chapter 1assessment Answers
 - The Rise of Digital Reading Prentice Hall Chemistry Chapter 1assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Chapter 1assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Chapter 1assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Chapter 1assessment Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Chapter 1assessment Answers User Reviews and Ratings
 - Prentice Hall Chemistry Chapter 1assessment Answers and Bestseller Lists

5. Accessing Prentice Hall Chemistry Chapter 1assessment Answers Free and Paid eBooks
 - Prentice Hall Chemistry Chapter 1assessment Answers Public Domain eBooks
 - Prentice Hall Chemistry Chapter 1assessment Answers eBook Subscription Services
 - Prentice Hall Chemistry Chapter 1assessment Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Chapter 1assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Chapter 1assessment Answers Compatibility with Devices
 - Prentice Hall Chemistry Chapter 1assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Chapter 1assessment Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Chapter 1assessment Answers
 - Interactive Elements Prentice Hall Chemistry Chapter 1assessment Answers
8. Staying Engaged with Prentice Hall Chemistry Chapter 1assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Chapter 1assessment Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Chapter 1assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Chapter 1assessment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Chapter 1assessment Answers
 - Setting Reading Goals Prentice Hall Chemistry Chapter 1assessment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Chapter 1assessment Answers
 - Fact-Checking eBook Content of Prentice Hall Chemistry Chapter 1assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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