

# Chemistry Preliminary Exam Notes

## Module 1 - Properties And Structure Of Matter

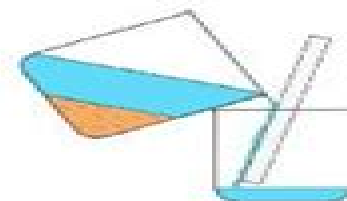
### How Do The Properties Of Substances Help Us To Classify And Separate Them?

#### WHATS IN THIS LESSON (PART 1 + PART 2):

- Explore homogeneous mixtures and heterogeneous mixtures through practical investigations:
  - Using separation techniques based off physical properties.
  - Calculating percentage composition by weight of component elements and/or compounds.

#### Separation Of Mixtures (Part 1):

- 2 Types Of Matter:
  - Pure Substance
  - Mixture
- Pure Substance
  - 2 Types
    - Elements
    - Compounds
- Mixtures
  - 2 Types
    - Homogeneous
    - Heterogeneous
  - Variable in composition
  - Can be separated by physical means.
    - E.G. Based on their size, base on temperatures they boil at.
- **Homogeneous:** A mixture with indistinguishable components.
  - E.G. Tea with sugar, air, tap water.
- **Heterogeneous:** A mixture where the components are physically distinguishable.
  - E.G. Oil and water, wet sand, bread, gravel, etc.
- **Phase:** State of Matter
  - E.G. Solid, liquid, gas, etc.
- Homogeneous Mixtures
  - Have the same consistency throughout.
  - All components are in (or dissolved into) a single phase/state of matter.
- Heterogeneous Mixtures
  - Differing consistency throughout
  - All components don't have to be in the same phase/state of matter.
- **Immiscible:** Two liquids that will not dissolve into one another. They do not form distinct layers.
- Separation Techniques
  - Decantation
    - **Theory Summary:** With immiscible liquids or mixtures with suspensions they will separate out according to density (with the most dense at the bottom and form layers. The top or bottom layer can be poured off - This is called decantation.
    - **Explanation:**
      - Physical property that separation is based on: density.
      - The mixture is allowed to settle (this may take longer with solids suspended in liquids).
      - One layer is poured off very carefully.
      - There are limitations to the purity of the separated components



# Preliminary Chemistry Revision

**Marcel A. Müller**



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## **Table of Contents Preliminary Chemistry Revision**

1. Understanding the eBook Preliminary Chemistry Revision
  - The Rise of Digital Reading Preliminary Chemistry Revision
  - Advantages of eBooks Over Traditional Books
2. Identifying Preliminary Chemistry Revision
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Preliminary Chemistry Revision
  - User-Friendly Interface
4. Exploring eBook Recommendations from Preliminary Chemistry Revision
  - Personalized Recommendations

- Preliminary Chemistry Revision User Reviews and Ratings
- Preliminary Chemistry Revision and Bestseller Lists
- 5. Accessing Preliminary Chemistry Revision Free and Paid eBooks
  - Preliminary Chemistry Revision Public Domain eBooks
  - Preliminary Chemistry Revision eBook Subscription Services
  - Preliminary Chemistry Revision Budget-Friendly Options
- 6. Navigating Preliminary Chemistry Revision eBook Formats
  - ePub, PDF, MOBI, and More
  - Preliminary Chemistry Revision Compatibility with Devices
  - Preliminary Chemistry Revision Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Preliminary Chemistry Revision
  - Highlighting and Note-Taking Preliminary Chemistry Revision
  - Interactive Elements Preliminary Chemistry Revision
- 8. Staying Engaged with Preliminary Chemistry Revision
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Preliminary Chemistry Revision
- 9. Balancing eBooks and Physical Books Preliminary Chemistry Revision
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Preliminary Chemistry Revision
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Preliminary Chemistry Revision
  - Setting Reading Goals Preliminary Chemistry Revision
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Preliminary Chemistry Revision
  - Fact-Checking eBook Content of Preliminary Chemistry Revision

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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