

Gas
Liquid
Solid
Unknown

Periodic Table

Cheat Sheet

Reactions

Single Displacement $\rightarrow A + BC = B + AC$
 Double Displacement $\rightarrow AB + CD = BC + AD$
 Synthesis $\rightarrow A + B = AB$
 Decomposition $\rightarrow AB = A + B$
 Combustion $\rightarrow \text{Hydrocarbon} + O_2 = CO_2 + H_2O$
 Neutralization $\rightarrow \text{Acid} + \text{Base} = \text{Salt} + H_2O$

Polyatomic ions (ion composed of multiple atoms)

$NO_3^- \rightarrow$ Nitrate (-1)

$PO_4^{3-} \rightarrow$ Phosphate (-3)

$OH^- \rightarrow$ Hydroxide (-1)

$HCO_3^- \rightarrow$ Bicarbonate (-1)

$NO_2^- \rightarrow$ Nitrite (-1)

$SO_4^{2-} \rightarrow$ Sulfate (-2)

$CO_3^{2-} \rightarrow$ Carbonate (-2)

$NH_4^+ \rightarrow$ Ammonium (+1)

H		Periodic Table										He					
Solid		Cheat Sheet										Unknown					
Polyatomic ions (ion composed of multiple atoms)																	
NO ₃ → Nitrate (-1)		NO ₂ → Nitrite (-1)															
PO ₄ → Phosphate (-3)		SO ₄ → Sulfate(-2)															
OH → Hydroxide (-1)		CO ₃ → Carbonate (-2)															
HCO ₃ → Bicarbonate (-1)		NH ₄ → Ammonium (+1)															
B		C		N		O		F		Ne							
Al		Si		P		S		Cl		Ar							
K		Ca		Sc		Ti		V		Cr		Mn		Fe		Co	
Rb		Sr		Y		Zr		Nb		Mo		Tc		Ru		Rh	
Cs		Ba		57-71		Hf		Ta		W		Re		Os		Ir	
Fr		Ra		89-103		Rf		Db		Sg		Bh		Hs		Mt	

Alkali Metals:
 One valence electron
 Very reactive because they are ready to lose an electron to form cations
 Usually in ionic bonds due to reactivity
 Bonds with halogens, and other non-metals
 Have a vigorous reaction when bonding with halogens, oxygen and the compound of water

Alkaline Earth Metals:
 Two valence electrons
 Inactive but not as reactive as halogens and alkali metals because it needs to lose 2 electrons to have a full valence shell
 Bonds with halogens, and other non-metals

Transition Metals/Metals: Transition metals are metals that vary in valence electrons and can have multiple ion charges. They have similar qualities as alkali metals and alkaline earth metals; they can form metallic and ionic bonds. Metals with multiple ion charges are called multivalent metals. For example, copper has an ion charge of +1 and +2 so it is a multivalent metal. Therefore, the ion charge must be specified when written in word form for ionic bonds with multivalent metals. Metals can also bond with polyatomic ions.

La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium	Lu Lutetium
Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium	Lr Lawrencium

Metalloids: Contain metal and non-metal properties. They react with metals and non-metals to form covalent and ionic bonds but are less reactive than non-metals, depending on the electronegativity of the atoms.

Non-metals: They can not conduct electricity because they can not give off electrons, so current will not be able to go through.
 Non-metals bond with metals in an ionic bond and other non-metals in a molecular bond. They can also bond with the polyatomic ion ammonium due to its +1 charge.

Noble Gases: Non-metals in a previous form that have a full valence shell. This makes it not reactive, so it can not bond with other non-metals and metals. This is why noble gases are never in compounds and chemical reactions.

Periodic Table Of Elements Grade 10

Didier Musso



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Table of Contents Periodic Table Of Elements Grade 10

1. Understanding the eBook Periodic Table Of Elements Grade 10
 - The Rise of Digital Reading Periodic Table Of Elements Grade 10
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Of Elements Grade 10
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Of Elements Grade 10
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Of Elements Grade 10
 - Personalized Recommendations
 - Periodic Table Of Elements Grade 10 User Reviews and Ratings
 - Periodic Table Of Elements Grade 10 and Bestseller Lists
5. Accessing Periodic Table Of Elements Grade 10 Free and Paid eBooks
 - Periodic Table Of Elements Grade 10 Public Domain eBooks
 - Periodic Table Of Elements Grade 10 eBook Subscription Services
 - Periodic Table Of Elements Grade 10 Budget-Friendly Options
6. Navigating Periodic Table Of Elements Grade 10 eBook Formats

- ePub, PDF, MOBI, and More
- Periodic Table Of Elements Grade 10 Compatibility with Devices
- Periodic Table Of Elements Grade 10 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Of Elements Grade 10
 - Highlighting and Note-Taking Periodic Table Of Elements Grade 10
 - Interactive Elements Periodic Table Of Elements Grade 10
- 8. Staying Engaged with Periodic Table Of Elements Grade 10
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Of Elements Grade 10
- 9. Balancing eBooks and Physical Books Periodic Table Of Elements Grade 10
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Of Elements Grade 10
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Of Elements Grade 10
 - Setting Reading Goals Periodic Table Of Elements Grade 10
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Of Elements Grade 10
 - Fact-Checking eBook Content of Periodic Table Of Elements Grade 10
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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