

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev stated that certain elements, under chemical properties of elements appeared at regular intervals which allowed him to
 - (a) classify
 - (b) determine
 - (c) determine
 - (d) determine
3. _____ The modern periodic law states that
 - (a) all new elements will be added upon as he listed in the table they were given
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a column
 - (c) adding a new group
 - (d) adding a new element
5. _____ The main reason the properties of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) unreactive
 - (d) large molecules
6. _____ Likewise, the first column in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 11
 - (c) 19
 - (d) 20
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 11
 - (b) 19
 - (c) 20
 - (d) 21
8. How many elements are in this group?
 - (a) 1
 - (b) 2
 - (c) 3
 - (d) 4
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 - (c) 3
 - (d) 4
100. How many elements are in this group?
 - (a) 1
 - (b) 2
 - (c) 3
 - (d) 4

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Michael Hoelscher

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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