

MYP4 Chemistry

Periodic Trends

Use the periodic table, charts, and your knowledge of periodic trends to answer the following questions.

Atomic Radius

1. Which atom in each pair has the larger atomic radius? Circle your answer.

a) Li or **K** b) **Ca** or Ni c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? From left to right in a period?

Atomic size (atomic radius) increases from top to bottom in a group and decreases from left to right in a period.

3. Why do atoms get smaller as you move left to right in a period?

Atoms get smaller as you move left to right in a period due to the increase in effective nuclear charge (while the number of occupied energy levels remain constant) leading to an increase in the force of attraction, which pulls the outer electrons closer to the nucleus and decreases the atomic radius.

4. Why do atoms get larger as you move down a group?

Atoms get larger as you move down a group due to the increase in occupied energy levels.

5. Arrange the following in order of increasing atomic radius atomic radius.

Mg, Na, P, Si and Ar? They are in the same period: Ar, P, Si, Mg, Na

Ionic Radius

6. Define an ion.

a charged atom or group of atoms

7. What is the difference between a cation and an anion?

Cations are positively charged ions. They are formed when an atom loses electrons.
Anions are negatively charged ions. They are formed when an atom gains electrons.

8. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Cations are smaller than their neutral atom because they have less occupied energy levels.
Anions are larger than their neutral atom because of the increase in electron-electron repulsion.

9. Which particle has the larger radius in each atom/ion pair? Circle your answer.

a) **Na** Na⁺ b) S, **S²⁻** c) I, I⁻ d) **Al** Al³⁺

Periodic Trends Practice Problems Answer Key

Michael Brown



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Table of Contents Periodic Trends Practice Problems Answer Key

1. Understanding the eBook Periodic Trends Practice Problems Answer Key
 - The Rise of Digital Reading Periodic Trends Practice Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Practice Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Practice Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Practice Problems Answer Key
 - Personalized Recommendations
 - Periodic Trends Practice Problems Answer Key User Reviews and Ratings
 - Periodic Trends Practice Problems Answer Key and Bestseller Lists

5. Accessing Periodic Trends Practice Problems Answer Key Free and Paid eBooks
 - Periodic Trends Practice Problems Answer Key Public Domain eBooks
 - Periodic Trends Practice Problems Answer Key eBook Subscription Services
 - Periodic Trends Practice Problems Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Practice Problems Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Practice Problems Answer Key Compatibility with Devices
 - Periodic Trends Practice Problems Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Practice Problems Answer Key
 - Highlighting and Note-Taking Periodic Trends Practice Problems Answer Key
 - Interactive Elements Periodic Trends Practice Problems Answer Key
8. Staying Engaged with Periodic Trends Practice Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Practice Problems Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Practice Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Practice Problems Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Practice Problems Answer Key
 - Setting Reading Goals Periodic Trends Practice Problems Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Practice Problems Answer Key
 - Fact-Checking eBook Content of Periodic Trends Practice Problems Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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