

Step 8. Periodic table trends

Electronegativity

Ionization energy

He

Atomic radius

The periodic table displays various elements with their symbols and atomic numbers. Four arrows indicate trends: a black arrow pointing right for electronegativity, a green arrow pointing up for ionization energy, a red arrow pointing down for atomic radius, and a blue arrow pointing from the bottom-left to the top-right for non-metallic properties. A light blue arrow points from the top-left to the bottom-right for metallic properties.

Ionization Energy

Electronegativity

Atomic radius

An arrow head shows the INCREASE of the property

Example: comparing K atom with Br atom

- Atomic radius of K is HIGHER than atomic radius of Br
- Ionization energy of K is LOWER than ionisation energy of Br
- Metallic properties of K are HIGHER than metallic properties of Br



Periodic Trends Answer Key Section Review

J Spring



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