

CHEMICAL BONDING

Type of bonds

- Chemical bond formation takes place when
(A) Energy is absorbed
(B) Forces of attraction are overcome by the forces of repulsion
(C) Forces of repulsion are overcome by forces of attraction
(D) Forces of attraction are equal to the forces of repulsion
- Most favourable conditions for ionic bond formation are
(A) Low charge on ions, large cation, small anion
(B) High charge on ions, large cation, small anion
(C) High charge on ions, small cation, large anion
(D) High charge on ions, small cation, large anion
- Which of the following shows the higher lattice energy?
(A) RbF (B) CsF (C) NaF (D) KF
- Which of the following is least soluble in water?
(A) BaF₂ (B) SiF₄ (C) CaF₂ (D) MgF₂
- Which of the following sets does not contain any ionic species:
(A) NH₄Cl, OF₂, H₂S (B) CO₂, Cl₂, CCl₄ (C) BF₃, AlF₃, TiF₃ (D) I₂, CaO, CH₃Cl
- There is no S-S bond in :
(A) S₂O₃²⁻ (B) S₂O₄²⁻ (C) S₄O₆²⁻ (D) S₂O₇²⁻
- The species that does not contain peroxide ions are :
(A) PbO₂ (B) H₂O₂ (C) SrO₂ (D) BaO₂
- Which of the following compounds contains electrovalent and covalent bonds?
(A) CS₂ (B) SO₂ (C) KCN (D) Phosphine
- In which solvent has maximum solubility
(A) H₂O (B) C₂H₅OH (C) CH₃COCH₃ (D) C₂H₅OC₂H₅
- Which of the following is a non-directional bond
(A) Metallic bond (B) Covalent bond (C) Co-ordinate bond (D) All are directional
- Find the number of σ and π bonds in the following compounds
(A) CH₃CH₃ (B) C₆(CN)₆ (C) CH₂=CH₂ (D) C₂H₂Cl
- Point out the inter particle force operating in each of the following :
(A) H₂O(l) (B) H₂O(s) (C) Dryice (D) KCl
(E) Hg (F) CF₄ (G) Rock salt (H) Diamond
- True / False
(i) Graphite, a covalent molecule, conducts current in solid state
(ii) Intra molecular H-Bonding decreases the solubility of compound in water

Matching

	Column I		Column - II
(a)	H ₂ O	(p)	Ionic bond
(b)	C ₂ H ₂	(q)	Covalent bond
(c)	KCN	(r)	Co-ordinate bond
(d)	K ₄ [Fe(CN) ₆]	(s)	Hydrogen bond

Fagan's rule, Physical property

- Amongst LiCl, BeCl₂, MgCl₂ and RbCl the compounds with greatest and least ionic character respectively are :
(A) LiCl and RbCl (B) RbCl and BeCl₂ (C) RbCl and MgCl₂ (D) MgCl₂ and BeCl₂
- The formal charge of the O atoms in the ion $[\text{O}=\text{N}=\text{O}]_{-2}$ is

Quiz On Chemical Bonding

David Kirk



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Table of Contents Quiz On Chemical Bonding

1. Understanding the eBook Quiz On Chemical Bonding
 - The Rise of Digital Reading Quiz On Chemical Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Quiz On Chemical Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Quiz On Chemical Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Quiz On Chemical Bonding
 - Personalized Recommendations
 - Quiz On Chemical Bonding User Reviews and Ratings
 - Quiz On Chemical Bonding and Bestseller Lists
5. Accessing Quiz On Chemical Bonding Free and Paid eBooks

- Quiz On Chemical Bonding Public Domain eBooks
- Quiz On Chemical Bonding eBook Subscription Services
- Quiz On Chemical Bonding Budget-Friendly Options
- 6. Navigating Quiz On Chemical Bonding eBook Formats
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 - Quiz On Chemical Bonding Compatibility with Devices
 - Quiz On Chemical Bonding Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Quiz On Chemical Bonding
 - Highlighting and Note-Taking Quiz On Chemical Bonding
 - Interactive Elements Quiz On Chemical Bonding
- 8. Staying Engaged with Quiz On Chemical Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Quiz On Chemical Bonding
- 9. Balancing eBooks and Physical Books Quiz On Chemical Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Quiz On Chemical Bonding
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Quiz On Chemical Bonding
 - Setting Reading Goals Quiz On Chemical Bonding
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Quiz On Chemical Bonding
 - Fact-Checking eBook Content of Quiz On Chemical Bonding
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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