

En un haluro de alquilo el átomo de halógeno está enlazado a un átomo de carbono con hibridación  $sp^3$ . El halógeno es más electronegativo que el carbono, y el enlace C—X está polarizado con una carga parcial positiva en el carbono y una carga parcial negativa en el halógeno.

## 6.4

### Estructura de los haluros de alquilo



$$\mu = 4.8 \times \delta \times d$$

donde  $\delta$  es la carga y  $d$  la longitud de enlace

Las electronegatividades de los halógenos *aumentan* en el siguiente orden:

|                    |     |   |     |   |     |   |     |
|--------------------|-----|---|-----|---|-----|---|-----|
|                    | I   | < | Br  | < | Cl  | < | F   |
| electronegatividad | 2.7 |   | 3.0 |   | 3.2 |   | 4.0 |

Las longitudes de enlace carbono-halógeno *aumentan* a medida que aumenta el volumen de los átomos de halógeno (mayor radio atómico) en el orden:

|                     |        |   |        |   |        |   |        |
|---------------------|--------|---|--------|---|--------|---|--------|
|                     | C—F    | < | C—Cl   | < | C—Br   | < | C—I    |
| longitud de enlace: | 1.38 Å |   | 1.78 Å |   | 1.94 Å |   | 2.14 Å |

Estos dos efectos son opuestos, ya que los halógenos que tienen más volumen y mayor longitud de enlace tienen las electronegatividades más pequeñas. El resultado total es que los momentos dipolares de enlace *aumentan* en el orden:

|                    |        |   |        |   |        |   |        |
|--------------------|--------|---|--------|---|--------|---|--------|
|                    | C—I    | < | C—Br   | < | C—F    | < | C—Cl   |
| momento dipolar, D | 1.29 D |   | 1.48 D |   | 1.51 D |   | 1.56 D |

# Quimicaorganica Webcindario Com

**Armando Novelli**



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