

Thermochemistry Chemistry Questions with Solutions

Q-1: Given that standard molar enthalpies of formation of NO(g) and $\text{NO}_2\text{(g)}$ are respectively 90.3 kJ/mol and 33.2 kJ/mol. Calculate the enthalpy change for the reaction $2\text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO}_2\text{(g)}$.

Answer: $\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$

$$\Delta H = 2(33.2) - [2(90.3) + 0]$$

$$\Delta H = -114.2 \text{ kJ}$$

Note: Enthalpy of formation of free elements like O_2 is always 0.

Q-2: The enthalpy of formation of AgCl is obtained from the enthalpy change from which one of the following process?

- a) $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl(s)}$
- b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$
- c) $\text{AgCl(s)} \rightarrow \text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g})$
- d) $\text{Ag(s)} + \text{AuCl} \rightarrow \text{Au(s)} + \text{AgCl(s)}$

Answer: b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Explanation: The enthalpy change accompanying the formation of one mole of a pure substance in its standard state from the constituent elements in their standard state is called enthalpy of formation. For Ag, the standard state is solid and for chlorine, it is gas. Thus, the process that gives the enthalpy of formation of AgCl is: $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Q-3: ΔH and ΔE for the reaction, $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(l)}$ at constant temperature are related as

- a) $\Delta H = \Delta E$
- b) $\Delta H = \Delta E + RT$
- c) $\Delta H = \Delta E + 3RT$
- d) $\Delta H = \Delta E - 3RT$

Answer: d) $\Delta H = \Delta E - 3RT$

Explanation: For any chemical reaction,

$$\Delta H = \Delta E + \Delta_{\text{ng}}RT$$

Where, Δ_{ng} = total number of moles of gaseous product- total number of moles of gaseous reactants.

Questions On Thermochemistry

JE Gale



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Table of Contents Questions On Thermochemistry

1. Understanding the eBook Questions On Thermochemistry
 - The Rise of Digital Reading Questions On Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Questions On Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Questions On Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Questions On Thermochemistry
 - Personalized Recommendations
 - Questions On Thermochemistry User Reviews and Ratings
 - Questions On Thermochemistry and Bestseller Lists
5. Accessing Questions On Thermochemistry Free and Paid eBooks
 - Questions On Thermochemistry Public Domain eBooks
 - Questions On Thermochemistry eBook Subscription Services
 - Questions On Thermochemistry Budget-Friendly Options

6. Navigating Questions On Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Questions On Thermochemistry Compatibility with Devices
 - Questions On Thermochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Questions On Thermochemistry
 - Highlighting and Note-Taking Questions On Thermochemistry
 - Interactive Elements Questions On Thermochemistry
8. Staying Engaged with Questions On Thermochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Questions On Thermochemistry
9. Balancing eBooks and Physical Books Questions On Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Questions On Thermochemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Questions On Thermochemistry
 - Setting Reading Goals Questions On Thermochemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Questions On Thermochemistry
 - Fact-Checking eBook Content of Questions On Thermochemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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