



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use blue or black ink for writing. Do not use red ink.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2 _____

QUESTION 3 starts on a NEW page.

- 3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 In what phase was compound A and B at 0 °C and 100 °C of boiling point? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What kind of the compound A and B is 1-hexanol? Explain how you arrived at your answer. (2)
- 3.1.5 What compound B and B₂ will have the higher vapour pressure? (2)
- 3.1.6 What compound B and B₂ will have the higher surface tension? (2)

Compound B (M_r 102) will show in the table has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is 16,2 °C. (2)

- 3.2 Add table 3.2 to table 3.1 and use the values of vapour enthalpies (ΔH_{vap}) and ΔH_{fusion} to calculate the values for each of the diagram below.



- 3.2.1 In which beaker (1) or (2), will the water evaporate? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the evaporation in the intermolecular forces involved. (2)

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