

## GRADE 12 PHYSICAL SCIENCES MARCH TEST

### QUESTION 1

Various options are provided as possible answers to the following questions. Choose the correct answer and write **ONLY** the LETTER (A-D) next to the QUESTION NUMBER (1.1-1.7) on your folio paper. Example 1.8. D.

- 1.1 Which of the following pairs of compounds contains members of the same homologous series?
- A Ethanol and butane  
 B Ethanol and propanoic acid  
 C Ethanol and propanal  
 D Ethanol and butanoic acid
- 1.2 In which of the following cases would the resultant force acting on an object be equal to 0N.



- A A constant number of different ions are released.  
 B A constant number of different ions are released, but the number of different ions is not constant.  
 C An infinite number of different ions are released, but the number of different ions is not constant.  
 D An infinite number of different ions are released, but the number of different ions is constant.
- 1.3 Which of the statements below best describes Weightlessness?
- A An object is free of weight because it is not in contact with the ground.  
 B An object is free of weight because it is not in contact with the ground.  
 C An object is free of weight because it is not in contact with the ground.  
 D An object is free of weight because it is not in contact with the ground.
- 1.4 The inertia of an object...
- A is the same for all objects.  
 B is the same for all objects.  
 C is the same for all objects.  
 D is the same for all objects.

- 1.5 An astronaut has a weight  $E_E$  on Earth. What distance, in terms of the radius of the Earth ( $r$ ), will the astronaut be if his weight is a 1/4 of his weight ( $E_E$ ) on the surface of the Earth?
- A  $r$   
 B  $2r$   
 C  $4r$   
 D  $8r$

- 1.6 A box is pushed forward with a force of 100N but does not move. A possible reason for this is that:
- A The box is too heavy.  
 B The box is too light.  
 C The box is too heavy.  
 D The box is too light.
- 1.7 A box is pushed forward with a force of 100N but does not move. A possible reason for this is that:
- A The box is too heavy.  
 B The box is too light.  
 C The box is too heavy.  
 D The box is too light.

[14]

### QUESTION 2

Three organic compounds (A, B and C) with different functional groups are given below.

A: 2,3-dimethyl-2-one



C:  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$

- 2.1 Write down the functional group of compound A. (1)  
 Write down the functional group of compound B. (1)  
 Write down the functional group of compound C. (1)  
 Write down the IUPAC name of compound B. (2)

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## **Table of Contents Physical Science Grade 1questions Paper March 2014**

1. Understanding the eBook Physical Science Grade 1questions Paper March 2014
  - The Rise of Digital Reading Physical Science Grade 1questions Paper March 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 1questions Paper March 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Grade 1questions Paper March 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 1questions Paper March 2014
  - Personalized Recommendations
  - Physical Science Grade 1questions Paper March 2014 User Reviews and Ratings
  - Physical Science Grade 1questions Paper March 2014 and Bestseller Lists

5. Accessing Physical Science Grade 1questions Paper March 2014 Free and Paid eBooks
  - Physical Science Grade 1questions Paper March 2014 Public Domain eBooks
  - Physical Science Grade 1questions Paper March 2014 eBook Subscription Services
  - Physical Science Grade 1questions Paper March 2014 Budget-Friendly Options
6. Navigating Physical Science Grade 1questions Paper March 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science Grade 1questions Paper March 2014 Compatibility with Devices
  - Physical Science Grade 1questions Paper March 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Grade 1questions Paper March 2014
  - Highlighting and Note-Taking Physical Science Grade 1questions Paper March 2014
  - Interactive Elements Physical Science Grade 1questions Paper March 2014
8. Staying Engaged with Physical Science Grade 1questions Paper March 2014
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Grade 1questions Paper March 2014
9. Balancing eBooks and Physical Books Physical Science Grade 1questions Paper March 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Grade 1questions Paper March 2014
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 1questions Paper March 2014
  - Setting Reading Goals Physical Science Grade 1questions Paper March 2014
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 1questions Paper March 2014
  - Fact-Checking eBook Content of Physical Science Grade 1questions Paper March 2014
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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$$m = \frac{\text{moles of solute}}{\text{kg of solvent}}$$

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web 1 molarity  $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$   $\frac{4500 \text{ g}}{0.825 \text{ g/mol} \times 0.0129 \text{ m}^3}$  2 molality  $m = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$   $\frac{0.825 \text{ g}}{141.9579 \text{ g/mol} \times 0.00581158 \text{ m}^3}$  3 mole fraction  $n_2 = \frac{\text{mass}}{\text{molar mass}}$   $\frac{4500 \text{ g}}{0.0129 \text{ m}^3}$  4  $\frac{0.825 \text{ g}}{141.9579 \text{ g/mol}}$  5  $\frac{0.00581158 \text{ m}^3}{4500 \text{ g}}$  6  $\frac{0.0129 \text{ m}^3}{141.9579 \text{ g/mol}}$  7  $\frac{4500 \text{ g}}{18.015 \text{ g/mol}}$  8  $\frac{4500 \text{ g}}{97918401 \text{ mol}}$

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web apr 7 2014 molarity moles of solute / liters of solution 8 4 2 2 a first convert 250 ml to liters  $\frac{250}{1000} = 0.25$  then calculate molarity  $\frac{5 \text{ moles}}{0.25 \text{ liters}} = 20 \text{ M}$  3 c a solution with molarity 2 requires 2 M of NaOH per liter so  $4 \times 2 = 8 \text{ M}$  4 a a

solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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$$M_{mix} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_{total}}$$
  
with increase in temperature which of these changes a molality b fraction of solute present in water

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