

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
$Mg(OH)_2$	<u>base</u>	H_3PO_4	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO_3	<u>salt</u>	$Al(OH)_3$	<u>base</u>
HFO_4	<u>acid</u>	$KC_2H_3O_2$	<u>salt</u>
$Ba(OH)_2$	<u>base</u>	NaCl	<u>salt</u>

Note Taking Worksheet Acids Bases And Salts

Brian J. Knapp



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Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Note Taking Worksheet Acids Bases And Salts**

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