

Periodic Trends Worksheet

Directions: Use the table to determine the following properties.

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Na, Al, S, O

2. Write the following elements by increasing electronegativity: sulfur, oxygen, sodium, aluminum.

O, S, Al, Na

3. Which atom has the higher first ionization energy: Na or K?

Na has a higher first ionization energy than K because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been two cases here where a strong periodic trend is present.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium, cesium, barium, thallium, lead, bismuth, polonium, astatine, and radon.

7. Which trend in ionization energy values across periods and the periodic table? Which element has the highest?

Ionization energy increases from left to right in periods and across the periodic table.

8. Write the atoms in each group that have the largest atomic radius.

a.	Li	Be
b.	Na	Mg
c.	Al	Si
d.	Ca	Sc
e.	K	Cr
f.	Rb	Ba

a.	Na	Ne
b.	Ar	Cl
c.	Br	Se
d.	As	Ge
e.	Sn	Pb

9. Circle the atom in each group that has the greatest electronegativity.

Physical Science If8767 Pg 63 Key

R Sanford

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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