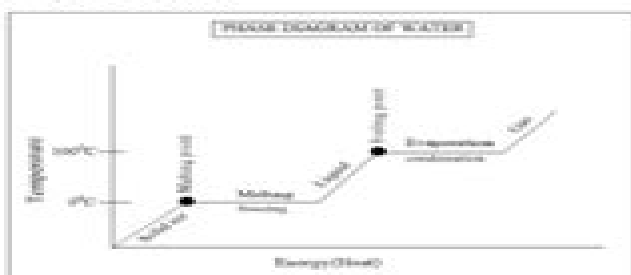


## Topic 1: Stoichiometry

### 1.1 Introduction to Matter & Chemical Change

- **matter:** has mass, consumes volume, made up of particles in constant motion
- **atomic theory:** all matter is composed of atoms that cannot be created nor destroyed
- **states of matter:**
  - **solid:** fixed volume/shape; cannot be compressed; attractive forces hold particles in close-packed arrangement; particles vibrate in fixed positions
  - **liquid:** fixed volume, no fixed shape; cannot be compressed; weaker forces; particles move around
  - **gas:** no fixed volume/shape; can be compressed
- as temperature ↑,  $KE_{\text{ave}}$  ↑
  - Kelvin (K): SI unit of temperature;  $^{\circ}\text{C} + 273$
- **changes of state:**



- **elements:** atoms of one type → form compounds
- **pure substance:** matter with constant composition → proportions are distinct
  - combine to form mixtures
  - homogeneous mixture: uniform composition and proportions
  - heterogeneous mixture: non-uniform composition and varying proportions
- **ion:** charged species
  - **anions:** negatively charged (contain excess electrons)
  - **cations:** positively charged (contains fewer electrons)
- **types of reactions:**
  - combination/synthesis rxns: combination of multiple substances
  - decomposition: breaking down of a single reactant
  - single displacement: one element replaces another in a compound
  - double displacement: swapping of elements

### 1.2 The Mole Concept

- **mole:** number of particles in a substance
  - fixed amount → Avogadro's number:  $6.02 \times 10^{23}$
- **isotopes:** different number of neutrons/mass number; same number of protons
  - relative abundance: measure of the percentage of isotopes in a sample
  - relative atomic mass ( $A_r$ ): weighted average of atomic masses of isotopes/relative abundance
    - relative molecular mass ( $M_r$ ): combination of all  $A_r$ s
- **molar mass:** mass of 1 mol of substance (g/mol)

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