

Parts of the Atom

**Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS
proton	proton	Nucleus	2000 X More than an e ⁻
neutron	neutron	Nucleus	2000 X More than an e ⁻
electron	electron	Orbitals / Energy Levels	1

**Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

6
C
carbon
12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element.

Error: Reference source not foundelement name: This is the name of the element.

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon:

Atomic mass: 12 (rounded)

Number of neutrons: $12 - 6 = 6$

To find the number of electrons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of electrons: 6

**Complete the table for the elements with atomic numbers 1-10.*

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

Note Taking Worksheet Inside The Atom Answers Packet

Daniel F McAuley



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Table of Contents Note Taking Worksheet Inside The Atom Answers Packet

1. Understanding the eBook Note Taking Worksheet Inside The Atom Answers Packet
 - The Rise of Digital Reading Note Taking Worksheet Inside The Atom Answers Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Worksheet Inside The Atom Answers Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Note Taking Worksheet Inside The Atom Answers Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Worksheet Inside The Atom Answers Packet
 - Personalized Recommendations
 - Note Taking Worksheet Inside The Atom Answers Packet User Reviews and Ratings
 - Note Taking Worksheet Inside The Atom Answers Packet and Bestseller Lists

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 - Note Taking Worksheet Inside The Atom Answers Packet Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Note Taking Worksheet Inside The Atom Answers Packet Compatibility with Devices
 - Note Taking Worksheet Inside The Atom Answers Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Worksheet Inside The Atom Answers Packet
 - Highlighting and Note-Taking Note Taking Worksheet Inside The Atom Answers Packet
 - Interactive Elements Note Taking Worksheet Inside The Atom Answers Packet
8. Staying Engaged with Note Taking Worksheet Inside The Atom Answers Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Worksheet Inside The Atom Answers Packet
9. Balancing eBooks and Physical Books Note Taking Worksheet Inside The Atom Answers Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Worksheet Inside The Atom Answers Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Worksheet Inside The Atom Answers Packet
 - Setting Reading Goals Note Taking Worksheet Inside The Atom Answers Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Worksheet Inside The Atom Answers Packet
 - Fact-Checking eBook Content of Note Taking Worksheet Inside The Atom Answers Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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