

Acid/Base indicator: \_\_\_\_\_ that changes \_\_\_\_\_ in the presence of an \_\_\_\_\_ or a \_\_\_\_\_.

| Solution                       | Litmus paper<br>(R→B, B→R,<br>or NC) | Phenolphthalein<br>(color or NC) | Bromothymol<br>Blue<br>(color or NC) | Cabbage Juice<br>(color or NC) |
|--------------------------------|--------------------------------------|----------------------------------|--------------------------------------|--------------------------------|
| HNO <sub>3</sub>               |                                      |                                  |                                      |                                |
| NaOH                           |                                      |                                  |                                      |                                |
| KOH                            |                                      |                                  |                                      |                                |
| H <sub>2</sub> SO <sub>4</sub> |                                      |                                  |                                      |                                |

Conclusion Questions:

1. Litmus turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
2. Phenolphthalein (phth) turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
3. Bromothymol blue turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.
4. Cabbage juice turns \_\_\_\_\_ in an acid and \_\_\_\_\_ in a base.

Strong Acids: dissociate \_\_\_\_\_ in \_\_\_\_\_ (ex: \_\_\_\_\_)  
 Weak Acids: dissociate \_\_\_\_\_ in \_\_\_\_\_  
 (ex: \_\_\_\_\_ or \_\_\_\_\_)

Strong Bases: dissociate \_\_\_\_\_ in \_\_\_\_\_ (ex: \_\_\_\_\_)  
 Weak Bases: dissociate \_\_\_\_\_ in \_\_\_\_\_  
 (ex: \_\_\_\_\_ or \_\_\_\_\_)

pH = \_\_\_\_\_

0 \_\_\_\_\_ 7 \_\_\_\_\_ 14

Determine the pH of a solution of HCl that has a molarity of  $1 \times 10^{-4} \text{ M}$ .

# Note Taking Guide Episode 1102 Answers

**L Cohen**



## **Note Taking Guide Episode 1102 Answers:**

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