

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING

The Ideal Gas Law

In 1811, the Italian chemist Amedeo Avogadro proposed the principle that *equal volumes of gases at the same temperature and pressure contain equal numbers of molecules*. He determined that at standard temperature and pressure, one mole of gas occupies 22.414 10 L (usually rounded to 22.4 L).

At this point, if you know the number of moles of a gas, you can use the molar volume of 22.4 L/mol to calculate the volume that amount of gas would occupy at STP. Then you could use the combined gas law to determine the volume of the gas under any other set of conditions. However, a much simpler way to accomplish the same task is by using the ideal gas law.

The *ideal gas law* is a mathematical relationship that has the conditions of standard temperature (273 K) and pressure (1 atm or 101.3 kPa) plus the molar gas volume (22.4 L/mol) already combined into a single constant. The following equation is the mathematical statement of the ideal gas law.

$$PV = nRT$$

in which

P = the pressure of a sample of gas

V = the volume of a sample of gas

n = the number of moles of gas present

T = the Kelvin temperature of the gas

R = the ideal gas constant, which combines standard conditions and molar volume into a single constant

The value of the ideal gas constant, R , depends on the units of P and V being used in the equation. Temperature is always in kelvins and amount of gas is always in moles. The most common values used for R are shown below.

Units of P and V	Value of R
Atmospheres and liters	$0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
Kilopascals and liters	$8.314 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$

Problem Solving Skills Chem File Mole

Daniela Niemeyer



Problem Solving Skills Chem File Mole:

Unveiling the Energy of Verbal Artistry: An Emotional Sojourn through **Problem Solving Skills Chem File Mole**

In a global inundated with monitors and the cacophony of instantaneous conversation, the profound energy and psychological resonance of verbal art often disappear in to obscurity, eclipsed by the regular onslaught of sound and distractions. However, nestled within the musical pages of **Problem Solving Skills Chem File Mole**, a interesting perform of literary brilliance that pulses with fresh feelings, lies an remarkable trip waiting to be embarked upon. Penned by a virtuoso wordsmith, that interesting opus books visitors on a mental odyssey, softly revealing the latent potential and profound impact stuck within the elaborate web of language. Within the heart-wrenching expanse of the evocative examination, we shall embark upon an introspective exploration of the book is central themes, dissect its captivating writing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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