

Chapter 13 Chemical Kinetics

Student: _____

1. The units of "reaction rate" are

A. $\text{L mol}^{-1} \text{s}^{-1}$
B. $\text{L}^2 \text{mol}^{-2} \text{s}^{-1}$
C. s^{-1}
D. s^{-2}
E. $\text{mol L}^{-1} \text{s}^{-1}$

2. For the reaction $\text{BrO}_3^- + 5\text{Br}^- + 6\text{H}^+ \rightarrow 3\text{Br}_2 + 3\text{H}_2\text{O}$ at a particular time, $-\Delta[\text{BrO}_3^-]/\Delta t = 1.5 \times 10^{-2} \text{ M/s}$. What is $-\Delta[\text{Br}^-]/\Delta t$ at the same instant?

A. 13 M/s
B. $7.5 \times 10^{-2} \text{ M/s}$
C. $1.5 \times 10^{-2} \text{ M/s}$
D. $3.0 \times 10^{-3} \text{ M/s}$
E. 330 M/s

3. For the following reaction, $\Delta P(\text{C}_6\text{H}_{14})/\Delta t$ was found to be $-6.2 \times 10^{-3} \text{ atm/s}$.
 $\text{C}_6\text{H}_{14}(\text{g}) \rightarrow \text{C}_6\text{H}_6(\text{g}) + 4\text{H}_2(\text{g})$

Determine $\Delta P(\text{H}_2)/\Delta t$ for this reaction at the same time.

A. $6.2 \times 10^{-3} \text{ atm/s}$
B. $1.6 \times 10^{-3} \text{ atm/s}$
C. $2.5 \times 10^{-2} \text{ atm/s}$
D. $-1.6 \times 10^{-3} \text{ atm/s}$
E. $-2.5 \times 10^{-2} \text{ atm/s}$

4. For the reaction $\text{C}_6\text{H}_{14}(\text{g}) \rightarrow \text{C}_6\text{H}_6(\text{g}) + 4\text{H}_2(\text{g})$, $\Delta P(\text{H}_2)/\Delta t$ was found to be $2.5 \times 10^{-2} \text{ atm/s}$, where $\Delta P(\text{H}_2)$ is the change in pressure of hydrogen. Determine $\Delta P(\text{C}_6\text{H}_{14})/\Delta t$ for this reaction at the same time.

A. $2.5 \times 10^{-2} \text{ atm/s}$
B. $-6.2 \times 10^{-3} \text{ atm/s}$
C. $-2.5 \times 10^{-2} \text{ atm/s}$
D. 0.10 atm/s
E. $6.2 \times 10^{-3} \text{ atm/s}$

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