

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{56}_{26}\text{Fe}$	B2																
3 (b)	$^{23}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th><th>charge</th></tr></thead><tbody><tr><td><math>^{17}\text{O}</math></td><td>8</td><td>9</td><td>10</td><td>2-</td></tr><tr><td><math>^{27}\text{Al}</math></td><td>13</td><td>14</td><td>10</td><td>3+</td></tr></tbody></table>		protons	neutrons	electrons	charge	$^{17}\text{O}$	8	9	10	2-	$^{27}\text{Al}$	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
$^{17}\text{O}$	8	9	10	2-														
$^{27}\text{Al}$	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"><thead><tr><th></th><th>protons</th><th>neutrons</th><th>electrons</th></tr></thead><tbody><tr><td><math>^{32}\text{S}</math></td><td>16</td><td>16</td><td>16</td></tr><tr><td><math>^{32}\text{S}^{2-}</math></td><td>16</td><td>16</td><td>18</td></tr></tbody></table>		protons	neutrons	electrons	$^{32}\text{S}$	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
$^{32}\text{S}$	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2(\text{SO}_4)_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

# Ocr Examination Answers As Exam Cafe Chemistry

**Tao Wei**



## **Ocr Examination Answers As Exam Cafe Chemistry:**

## **Ocr Examination Answers As Exam Cafe Chemistry** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Ocr Examination Answers As Exam Cafe Chemistry**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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