

- 99) Ionization energies allow chemists to determine electron structures. The first two ionization energies of rubidium and strontium are given in the table.

| element | first ionization energy (kJ/mol) | second ionization energy (kJ/mol) |
|-----------|----------------------------------|-----------------------------------|
| rubidium | 403 | 2633 |
| strontium | 550 | 1064 |

- 100) Write an equation to represent the **second** ionization energy of strontium.

Include state symbols.

[2]

- 101) Why is the **first** ionization energy of strontium larger than the **first** ionization energy of rubidium?

In your answer you should use appropriate chemical terms **correctly**.

[2]

- 102) Why is the **second** ionization energy of rubidium larger than the **second** ionization energy of strontium?

[2]

(Total 12)

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Jianjun Gao



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