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web 33mol o<sub>2</sub> 6mol br<sub>2</sub> 11mol o<sub>2</sub> 18mol br<sub>2</sub> identify the limiting reactant s and excess reactant s the limiting reactant is o<sub>2</sub>

since it would yield the least amount of product 18 mol br 2 the excess reactant is c 2 h 3 br 3 since its complete reaction would have yielded up to 54 mol br 2 think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o2 o 2 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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