

F321 Revision of definitions to be learnt for exam

Isotopes	Atoms of the same element with different numbers of neutrons and different masses.
Relative atomic mass, A_r	The weighted mean mass of an atom of an element compared with one-twelfth of the mass of an atom of ^{12}C .
Relative isotopic mass	The mass of an atom of an isotope compared with one-twelfth of the mass of an atom of carbon-12.
Relative formula mass	The relative formula mass of a compound is equal to the sum of the individual relative atomic masses.
Relative molecular mass, M_r	The weighted mean mass of a molecule compared with one-twelfth of the mass of an atom of carbon-12.
Mole	The amount of any substance containing as many particles as there are carbon atoms in exactly 12 g of the carbon-12 isotope.
Molecular formula	The number of atoms of each element in a molecule.
Empirical formula	The simplest whole-number ratio of atoms of each element present in a compound.
Acid	A species that is a proton donor.
Base	A species that is a proton acceptor.
Concentrated Solution	Solutions containing a higher number of moles of solute in a given volume of solution
Dilute Solution	Solutions containing a fewer number of moles of solute in a given volume of solution
Hydrated	Salt containing water molecules known as "water of crystallisation"
Anhydrous	A substance that contains no water of crystallisation
Oxidation	Loss of electrons, or an increase in oxidation number.
Reduction	Gain of electrons, or a decrease in oxidation number.
Redox reaction	A reaction in which both reduction and oxidation take place.
(First) ionisation energy	The energy required to remove one electron from each atom in one mole of gaseous atoms to form one mole of gaseous $1+$ ions.
(Second) ionisation energy	The energy required to remove one electron from each ion in one mole of gaseous $1+$ ions to form one mole of gaseous $2+$ ions.

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David Baud



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Table of Contents Ocr Chemistry 2013 January Past Paper F321

1. Understanding the eBook Ocr Chemistry 2013 January Past Paper F321
 - The Rise of Digital Reading Ocr Chemistry 2013 January Past Paper F321
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry 2013 January Past Paper F321
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry 2013 January Past Paper F321
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry 2013 January Past Paper F321
 - Personalized Recommendations
 - Ocr Chemistry 2013 January Past Paper F321 User Reviews and Ratings
 - Ocr Chemistry 2013 January Past Paper F321 and Bestseller Lists
5. Accessing Ocr Chemistry 2013 January Past Paper F321 Free and Paid eBooks
 - Ocr Chemistry 2013 January Past Paper F321 Public Domain eBooks
 - Ocr Chemistry 2013 January Past Paper F321 eBook Subscription Services
 - Ocr Chemistry 2013 January Past Paper F321 Budget-Friendly Options
6. Navigating Ocr Chemistry 2013 January Past Paper F321 eBook Formats

- ePub, PDF, MOBI, and More
- Ocr Chemistry 2013 January Past Paper F321 Compatibility with Devices
- Ocr Chemistry 2013 January Past Paper F321 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry 2013 January Past Paper F321
 - Highlighting and Note-Taking Ocr Chemistry 2013 January Past Paper F321
 - Interactive Elements Ocr Chemistry 2013 January Past Paper F321
- 8. Staying Engaged with Ocr Chemistry 2013 January Past Paper F321
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry 2013 January Past Paper F321
- 9. Balancing eBooks and Physical Books Ocr Chemistry 2013 January Past Paper F321
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry 2013 January Past Paper F321
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry 2013 January Past Paper F321
 - Setting Reading Goals Ocr Chemistry 2013 January Past Paper F321
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry 2013 January Past Paper F321
 - Fact-Checking eBook Content of Ocr Chemistry 2013 January Past Paper F321
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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