

## Reaction Rates Test Name:

\_\_\_\_K \_\_\_\_T \_\_\_\_C \_\_\_\_A  
20 9 9 8

## Assessment of Learning (Product)

Multiple Choice; Circle the most correct answer. (20K)

1. The decomposition of  $\text{N}_2\text{O}_5$  in solution in carbon tetrachloride proceeds via the reaction



The reaction is first order and has a rate constant of  $4.82 \times 10^{-5} \text{ s}^{-1}$  at  $64^\circ\text{C}$ . The rate law for the reaction is rate = \_\_\_\_\_.

- A)  $k[\text{N}_2\text{O}_5]$     B)  $k \frac{[\text{NO}_2][\text{O}_2]}{[\text{N}_2\text{O}_5]}$     C)  $k[\text{N}_2\text{O}_5]^2$     D)  $k \frac{[\text{N}_2\text{O}_5]}{[\text{NO}_2][\text{O}_2]}$     E)  $2k[\text{N}_2\text{O}_5]$

2. A burning splint will burn more vigorously in pure oxygen than in air because:

- A) oxygen is a reactant in combustion and concentration of oxygen is higher in pure oxygen than is in air.  
B) oxygen is a catalyst for combustion.  
C) oxygen is a product of combustion.  
D) nitrogen is a product of combustion and the system reaches equilibrium at a lower temperature  
E) nitrogen is a reactant in combustion and its low concentration in pure oxygen catalyzes the combustion.

3. Of the following, all are valid units for a reaction rate except \_\_\_\_\_.

- A)  $\text{mol/L}$     B)  $\text{M/s}$     C)  $\text{mol/hr}$     D)  $\text{s}^{-1}$     E)  $\text{mol/L}\cdot\text{h}$

4. Which of the following changes occur when the temperature of a reaction is increased?

- I.  $\Delta H$  of the reaction increases  
II. Frequency of the collisions increases  
III. Kinetic energy of the reactants increases

- A) I and II only.    B) I and III only.    C) II and III only.    D) I, II and III.

5. Which of the following is true for an activated complex?

- A) stable and has low PE    C) unstable and has low PE  
B) stable and has high PE    D) unstable and has high PE

6. Consider the following reaction:



The activation energy for the formation of HI is 167 kJ. The activation energy for the decomposition of HI is:

- A) 28 kJ    B) 139 kJ    C) 167 kJ    D) 195 kJ

# Reaction Rates Chemistry Packet Answers

**M Lipman**



## **Reaction Rates Chemistry Packet Answers:**

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### **Table of Contents Reaction Rates Chemistry Packet Answers**

1. Understanding the eBook Reaction Rates Chemistry Packet Answers
  - The Rise of Digital Reading Reaction Rates Chemistry Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates Chemistry Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Rates Chemistry Packet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates Chemistry Packet Answers
  - Personalized Recommendations

- Reaction Rates Chemistry Packet Answers User Reviews and Ratings
- Reaction Rates Chemistry Packet Answers and Bestseller Lists
- 5. Accessing Reaction Rates Chemistry Packet Answers Free and Paid eBooks
  - Reaction Rates Chemistry Packet Answers Public Domain eBooks
  - Reaction Rates Chemistry Packet Answers eBook Subscription Services
  - Reaction Rates Chemistry Packet Answers Budget-Friendly Options
- 6. Navigating Reaction Rates Chemistry Packet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Reaction Rates Chemistry Packet Answers Compatibility with Devices
  - Reaction Rates Chemistry Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reaction Rates Chemistry Packet Answers
  - Highlighting and Note-Taking Reaction Rates Chemistry Packet Answers
  - Interactive Elements Reaction Rates Chemistry Packet Answers
- 8. Staying Engaged with Reaction Rates Chemistry Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reaction Rates Chemistry Packet Answers
- 9. Balancing eBooks and Physical Books Reaction Rates Chemistry Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reaction Rates Chemistry Packet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rates Chemistry Packet Answers
  - Setting Reading Goals Reaction Rates Chemistry Packet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rates Chemistry Packet Answers
  - Fact-Checking eBook Content of Reaction Rates Chemistry Packet Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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