

$$r^3 \propto A$$

$$r \propto A^{1/3}$$

- (b) Using the equation given in the problem, we can first solve for the radius of the lithium nucleus and then solve for its volume.

$$r = r_0 A^{1/3}$$

$$r = (1.2 \times 10^{-15} \text{ m})(7)^{1/3}$$

$$r = 2.3 \times 10^{-15} \text{ m}$$

$$V = \frac{4}{3} \pi r^3$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi (2.3 \times 10^{-15} \text{ m})^3 = 5.1 \times 10^{-44} \text{ m}^3$$

- (c) In part (b), the volume of the nucleus was calculated. Using the radius of a Li atom, the volume of a Li atom can be calculated.

$$V_{\text{atom}} = \frac{4}{3} \pi r^3 = \frac{4}{3} \pi (152 \times 10^{-12} \text{ m})^3 = 1.47 \times 10^{-29} \text{ m}^3$$

The fraction of the atom's volume occupied by the nucleus is:

$$\frac{V_{\text{nucleus}}}{V_{\text{atom}}} = \frac{5.1 \times 10^{-44} \text{ m}^3}{1.47 \times 10^{-29} \text{ m}^3} = 3.5 \times 10^{-15}$$

Yes, this calculation shows that the volume of the nucleus is *much, much* smaller than the volume of the atom, which supports Rutherford's model of an atom.

- 2.99** Two different structural formulas for the molecular formula $\text{C}_2\text{H}_6\text{O}$ are:



In the second hypothesis of Dalton's Atomic Theory, he states that in any compound, the ratio of the number of atoms of any two of the elements present is either an integer or simple fraction. In the above two compounds, the ratio of atoms is the same. This does not necessarily contradict Dalton's hypothesis, but Dalton was not aware of chemical bond formation and structural formulas.

- 2.100** (a)

Ethane	Acetylene
2.65 g C	4.56 g C
0.665 g H	0.383 g H

Let's compare the ratio of the hydrogen masses in the two compounds. To do this, we need to start with the same mass of carbon. If we were to start with 4.56 g of C in ethane, how much hydrogen would combine with 4.56 g of carbon?

$$0.665 \text{ g H} \times \frac{4.56 \text{ g C}}{2.65 \text{ g C}} = 1.14 \text{ g H}$$

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Terry C. Jones



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Table of Contents Raymond Chang 10th Edition Solution Manual

1. Understanding the eBook Raymond Chang 10th Edition Solution Manual
 - The Rise of Digital Reading Raymond Chang 10th Edition Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Raymond Chang 10th Edition Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Raymond Chang 10th Edition Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Raymond Chang 10th Edition Solution Manual
 - Personalized Recommendations
 - Raymond Chang 10th Edition Solution Manual User Reviews and Ratings

- Raymond Chang 10th Edition Solution Manual and Bestseller Lists
- 5. Accessing Raymond Chang 10th Edition Solution Manual Free and Paid eBooks
 - Raymond Chang 10th Edition Solution Manual Public Domain eBooks
 - Raymond Chang 10th Edition Solution Manual eBook Subscription Services
 - Raymond Chang 10th Edition Solution Manual Budget-Friendly Options
- 6. Navigating Raymond Chang 10th Edition Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Raymond Chang 10th Edition Solution Manual Compatibility with Devices
 - Raymond Chang 10th Edition Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Raymond Chang 10th Edition Solution Manual
 - Highlighting and Note-Taking Raymond Chang 10th Edition Solution Manual
 - Interactive Elements Raymond Chang 10th Edition Solution Manual
- 8. Staying Engaged with Raymond Chang 10th Edition Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Raymond Chang 10th Edition Solution Manual
- 9. Balancing eBooks and Physical Books Raymond Chang 10th Edition Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Raymond Chang 10th Edition Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Raymond Chang 10th Edition Solution Manual
 - Setting Reading Goals Raymond Chang 10th Edition Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Raymond Chang 10th Edition Solution Manual
 - Fact-Checking eBook Content of Raymond Chang 10th Edition Solution Manual
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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