

OCR A LEVEL CHEMISTRY DEFINITIONS

Module 2:	
Term	Definition
Isotope	atoms of the same element with different numbers of neutrons and different masses
Relative isotopic mass	Mass compared with 1/12th mass of carbon-12
Relative atomic mass	weighted mean mass compared with 1/12th mass of carbon-12 atom.
Empirical formula	the simplest whole number ratio of atoms of each element present in a compound
Molecular formula	the number and type of atoms of each element in a molecule
Acid/Bronsted Lowry Acid	release H ⁺ ions in aqueous solution/Proton donor
Alkalis/ Bronsted Lowry Base	release OH ⁻ ions in aqueous solution/proton acceptor
Strong acid	Fully dissociates to release H ⁺ ions in aqueous solution
Weak acid	Partially dissociates to release H ⁺ ions in aqueous solution
Orbital	a region around the nucleus that can hold up to two electrons, with opposite spins
Ionic Bond	electrostatic attraction between positive and negative ions
Covalent Bond	strong electrostatic attraction between a shared pair of electrons and the nuclei of the bonded atoms
Metallic Bond	strong electrostatic attraction between cations (positive ions) and delocalised electrons
Electronegativity	ability of an atom to attract the bonding electrons in a covalent bond
Module 3:	
First Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous atoms to form 1 mol of gaseous 1+ ions
Second Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous 1+ ions to form 1 mol of gaseous 2+ ions
Disproportionation	oxidation and reduction of the same element
Standard conditions	298K, 1atm/100kPa, 1 moldm ⁻³ solutions
Standard enthalpy of formation	the enthalpy change when 1 mole of a compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion	the enthalpy change when 1 mole of a substance is completely burned in oxygen under standard conditions
Standard enthalpy of neutralisation	the enthalpy change when 1 mole of an acid reacts with 1 mole of an alkali to form 1 mole of water under standard conditions
Standard enthalpy of solution	the enthalpy change when 1 mole of an ionic substance dissolves in water to form 1 mole of aqueous ions under standard conditions
Standard enthalpy of hydration	the enthalpy change when 1 mole of gaseous ions is hydrated under standard conditions
Standard enthalpy of atomisation	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions
Standard enthalpy of formation of an ionic compound	the enthalpy change when 1 mole of an ionic compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion of an ionic compound	the enthalpy change when 1 mole of an ionic compound is completely burned in oxygen under standard conditions
Standard enthalpy of neutralisation of an ionic compound	the enthalpy change when 1 mole of an ionic compound reacts with 1 mole of an alkali to form 1 mole of water under standard conditions
Standard enthalpy of solution of an ionic compound	the enthalpy change when 1 mole of an ionic compound dissolves in water to form 1 mole of aqueous ions under standard conditions
Standard enthalpy of hydration of an ionic compound	the enthalpy change when 1 mole of gaseous ions is hydrated under standard conditions
Standard enthalpy of atomisation of an ionic compound	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions
Standard enthalpy of formation of a covalent compound	the enthalpy change when 1 mole of a covalent compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion of a covalent compound	the enthalpy change when 1 mole of a covalent compound is completely burned in oxygen under standard conditions
Standard enthalpy of neutralisation of a covalent compound	the enthalpy change when 1 mole of a covalent compound reacts with 1 mole of an alkali to form 1 mole of water under standard conditions
Standard enthalpy of solution of a covalent compound	the enthalpy change when 1 mole of a covalent compound dissolves in water to form 1 mole of aqueous ions under standard conditions
Standard enthalpy of hydration of a covalent compound	the enthalpy change when 1 mole of gaseous ions is hydrated under standard conditions
Standard enthalpy of atomisation of a covalent compound	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions

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Judd E. Hollander



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Table of Contents Ocr Alevel F32past Papers 23 May

1. Understanding the eBook Ocr Alevel F32past Papers 23 May
 - The Rise of Digital Reading Ocr Alevel F32past Papers 23 May
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Alevel F32past Papers 23 May
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Alevel F32past Papers 23 May
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Alevel F32past Papers 23 May
 - Personalized Recommendations
 - Ocr Alevel F32past Papers 23 May User Reviews and Ratings
 - Ocr Alevel F32past Papers 23 May and Bestseller Lists

5. Accessing Ocr Alevel F32past Papers 23 May Free and Paid eBooks
 - Ocr Alevel F32past Papers 23 May Public Domain eBooks
 - Ocr Alevel F32past Papers 23 May eBook Subscription Services
 - Ocr Alevel F32past Papers 23 May Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Ocr Alevel F32past Papers 23 May Compatibility with Devices
 - Ocr Alevel F32past Papers 23 May Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Alevel F32past Papers 23 May
 - Highlighting and Note-Taking Ocr Alevel F32past Papers 23 May
 - Interactive Elements Ocr Alevel F32past Papers 23 May
8. Staying Engaged with Ocr Alevel F32past Papers 23 May
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Alevel F32past Papers 23 May
9. Balancing eBooks and Physical Books Ocr Alevel F32past Papers 23 May
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Alevel F32past Papers 23 May
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Alevel F32past Papers 23 May
 - Setting Reading Goals Ocr Alevel F32past Papers 23 May
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Alevel F32past Papers 23 May
 - Fact-Checking eBook Content of Ocr Alevel F32past Papers 23 May
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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