

Lesson Quiz

Observing Chemical Change

Write the letter of the correct answer on the line at the left.

1. ____ Which of the following is true about chemical reactions?
 - A They are accompanied by changes in energy.
 - B They form new substances with new properties.
 - C both A and B
 - D neither A nor B
2. ____ In an endothermic reaction, energy is
 - A absorbed
 - B released
 - C converted to mass
 - D synthesized
3. ____ Which of the following is NOT a physical property?
 - A melting point
 - B state of matter
 - C density
 - D flammability
4. ____ Substances formed as a result of a chemical reaction are called
 - A catalysts
 - B precipitates
 - C products
 - D reactants

If the statement is true, write *true*. If the statement is false, change the underlined word or words to make the statement true.

5. _____ In an exothermic reaction, products have more energy than reactants.
6. _____ Water boils at 100°C. This is an example of a chemical property.
7. _____ Substances that enter into a chemical reaction are called products.
8. _____ The ability to react with oxygen is an example of a chemical property.
9. _____ Another name for a chemical change is a chemical bond.
10. _____ In a physical change, some of the physical properties of the substance may be altered and the chemical composition remains the same.

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A Loxley



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Pearson Edexcel A Level Chemistry (Year 1 and Year 2) Andrew Hunt, Graham Curtis, Graham Hill, 2019-07-15 Develop and assess your students knowledge and skills throughout A level with worked examples practical assessment guidance and differentiated end of topic questions in this updated all in one textbook for Years 1 and 2 Combining everything your students need to know for the Pearson Edexcel A level Chemistry specification this revised textbook will Identify the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Student Book Provide support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Improve mathematical skills with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offer plenty of practice with Test yourself questions to help students assess their understanding and measure progress Encourage further reading and study with short passages of extension material Develop understanding with free online access to Test yourself answers and an extended glossary **Focus on Physical Science California Edition** Michael J. Padilla, 2007 *FCS physical science L2*, 2007

Introductory Chemistry Charles H. Corwin, 2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking

Pharmaceutical Journal; ,1908 **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction The Pharmaceutical Journal and Pharmacist ,1918 **Insights into Chemical Engineering** P. V. Danckwerts, 2013-10-22 A selection of papers many of which proved novel and thought provoking and have had a considerable influence on the development of chemical engineering chosen by Professor Danckwerts from research work conducted at Cambridge and Imperial College mainly during the years 1950 1954 and 1957 1973 They are divided into 6 sections with linking critical commentaries

Geomorphology Robert S. Anderson, Suzanne P. Anderson, 2010-06-17 This textbook provides a modern quantitative and process oriented approach to equip students with the tools to understand geomorphology Insight into the interpretation of landscapes is developed from basic principles and simple models and by stepping through the equations that capture the essence of the mechanics and chemistry of landscapes Boxed worked examples and real world applications bring the subject

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